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**Hays et al.**

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(54) **NANOSTRUCTURED PLATINUM ALLOYS  
FOR USE AS CATALYST MATERIALS**

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May 19, 2010.

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**C23C 14/32** (2006.01)

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502/527.2; 204/192.15; 977/773-777;  
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See application file for complete search history.

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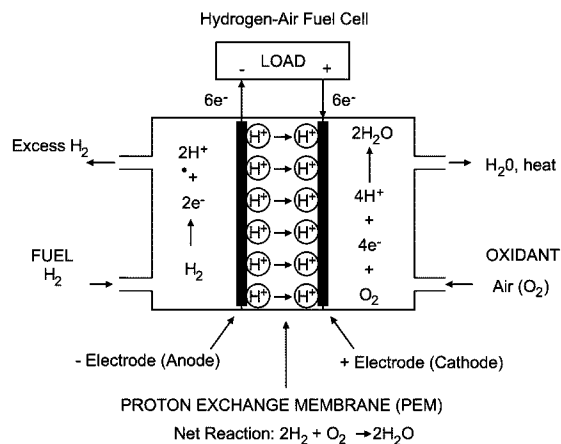
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(57) **ABSTRACT**

A series of binary and ternary Pt-alloys, that promote the  
important reactions for catalysis at an alloy surface; oxygen  
reduction, hydrogen oxidation, and hydrogen and oxygen  
evolution. The first two of these reactions are essential when  
applying the alloy for use in a PEMFC.

**27 Claims, 21 Drawing Sheets**



Hydrogen-Air fuel cell schematic

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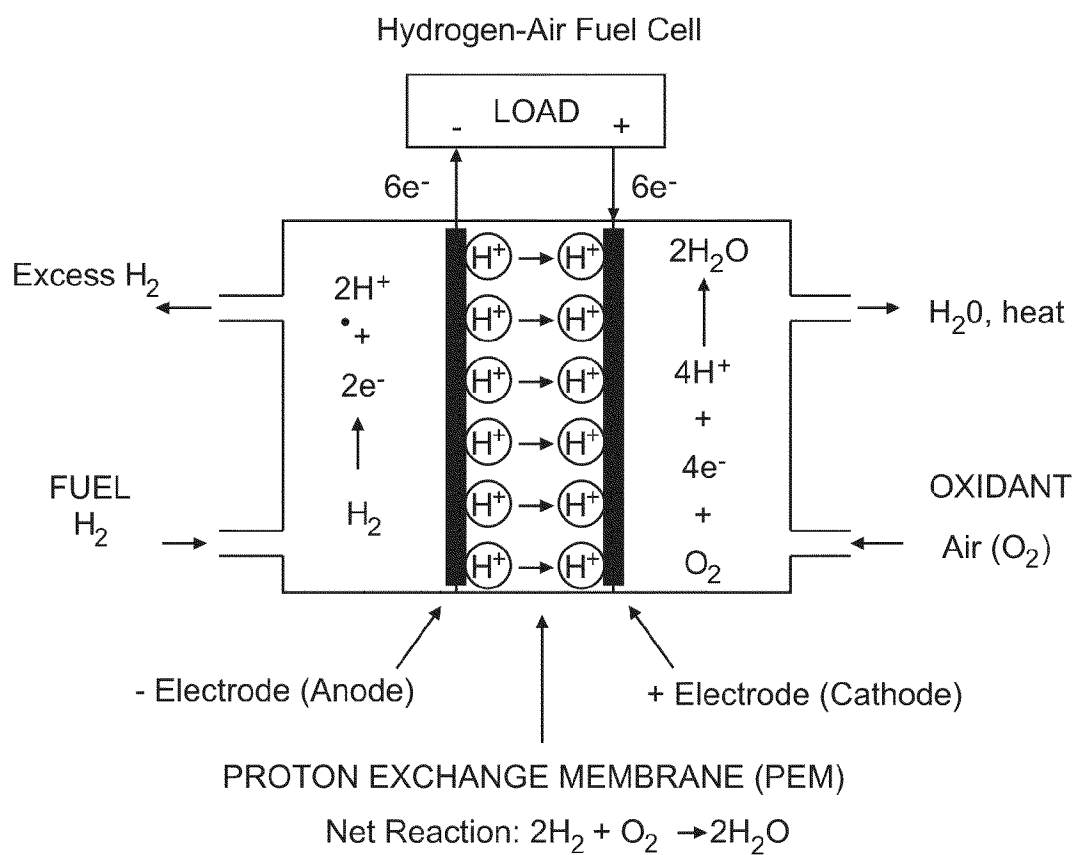


Figure 1: Hydrogen-Air fuel cell schematic

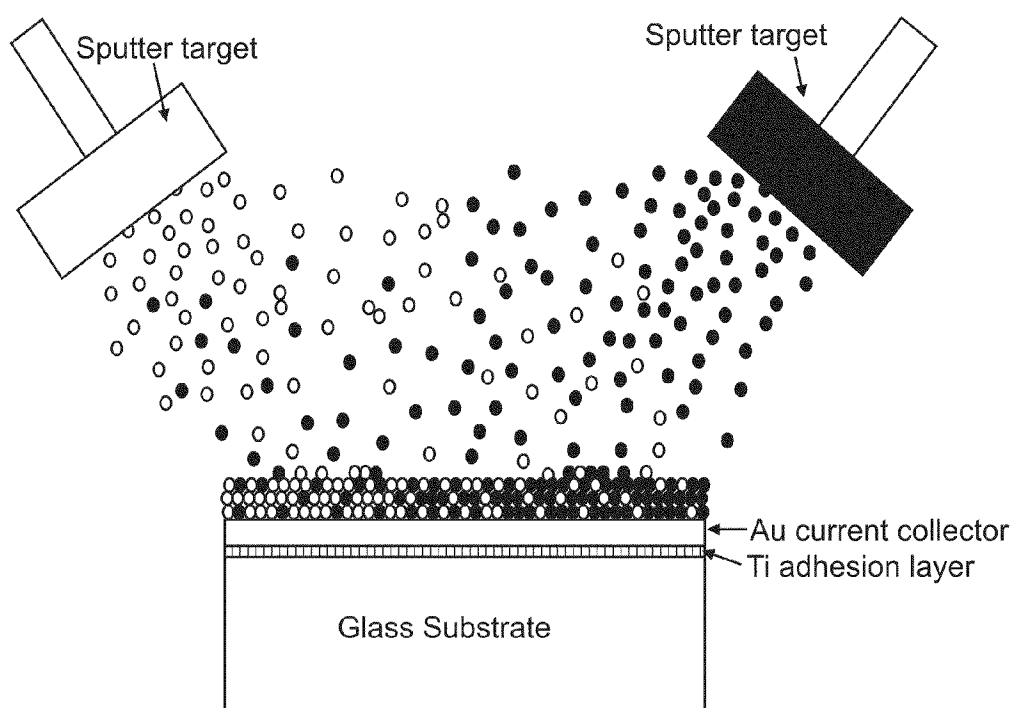


Figure 2a: Co-sputtering

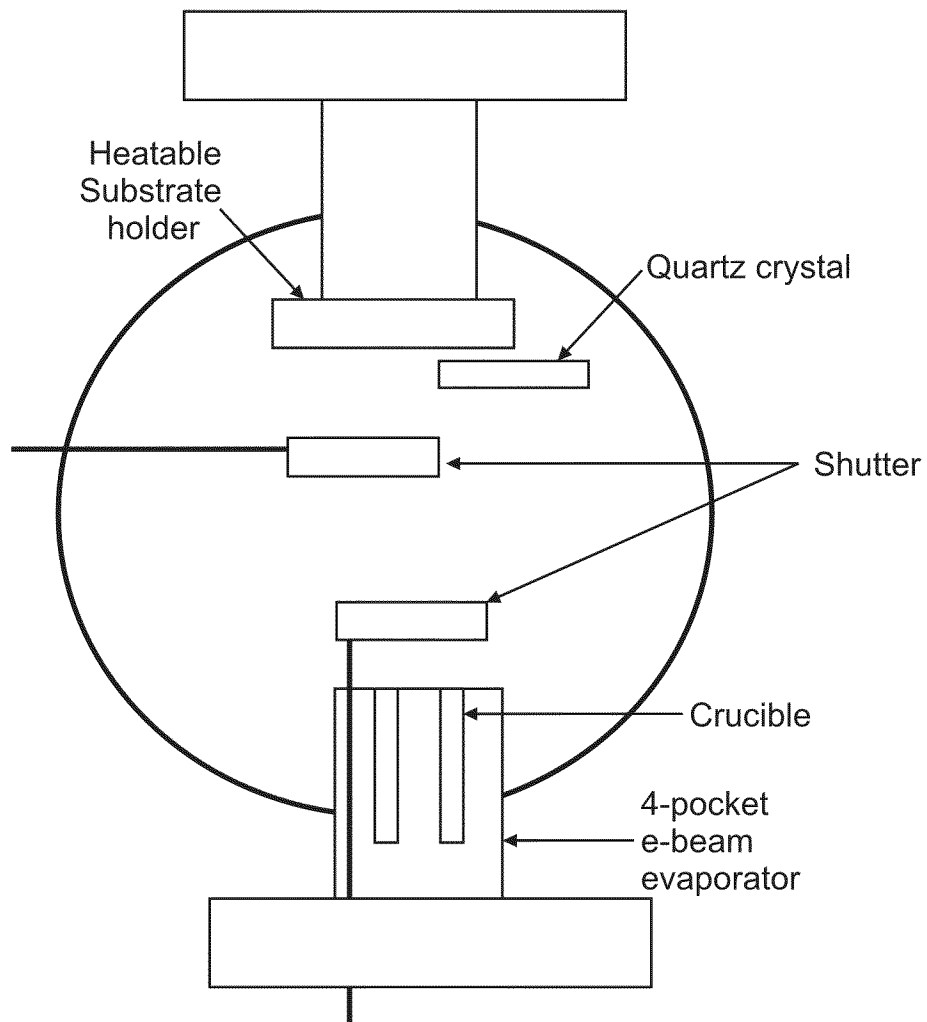


Figure 2b: Electron Beam Evaporation

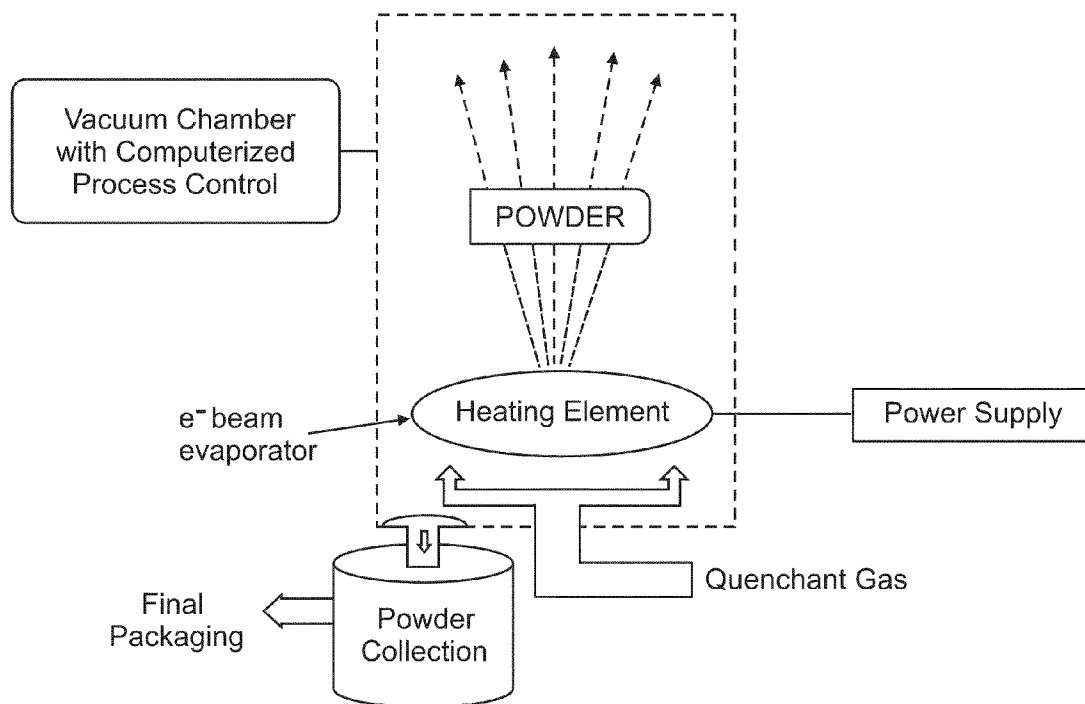
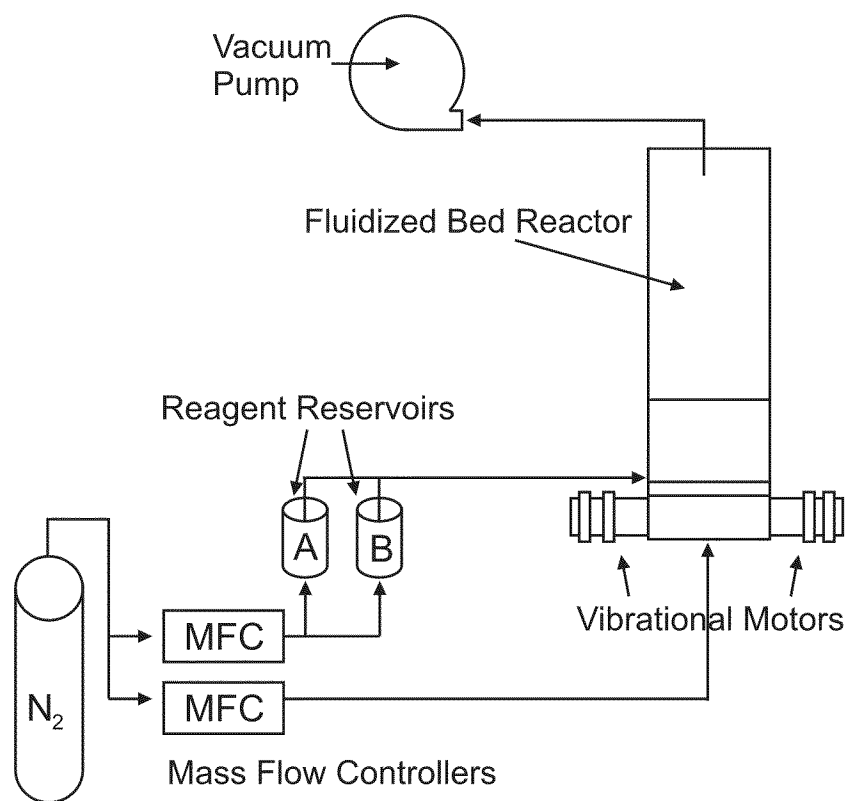


Figure 2c: Vapor-phase condensation



(F) Fluidized bed ALD reactor schematic

Figure 2d: Atomic-Layer-Deposition (ALD)

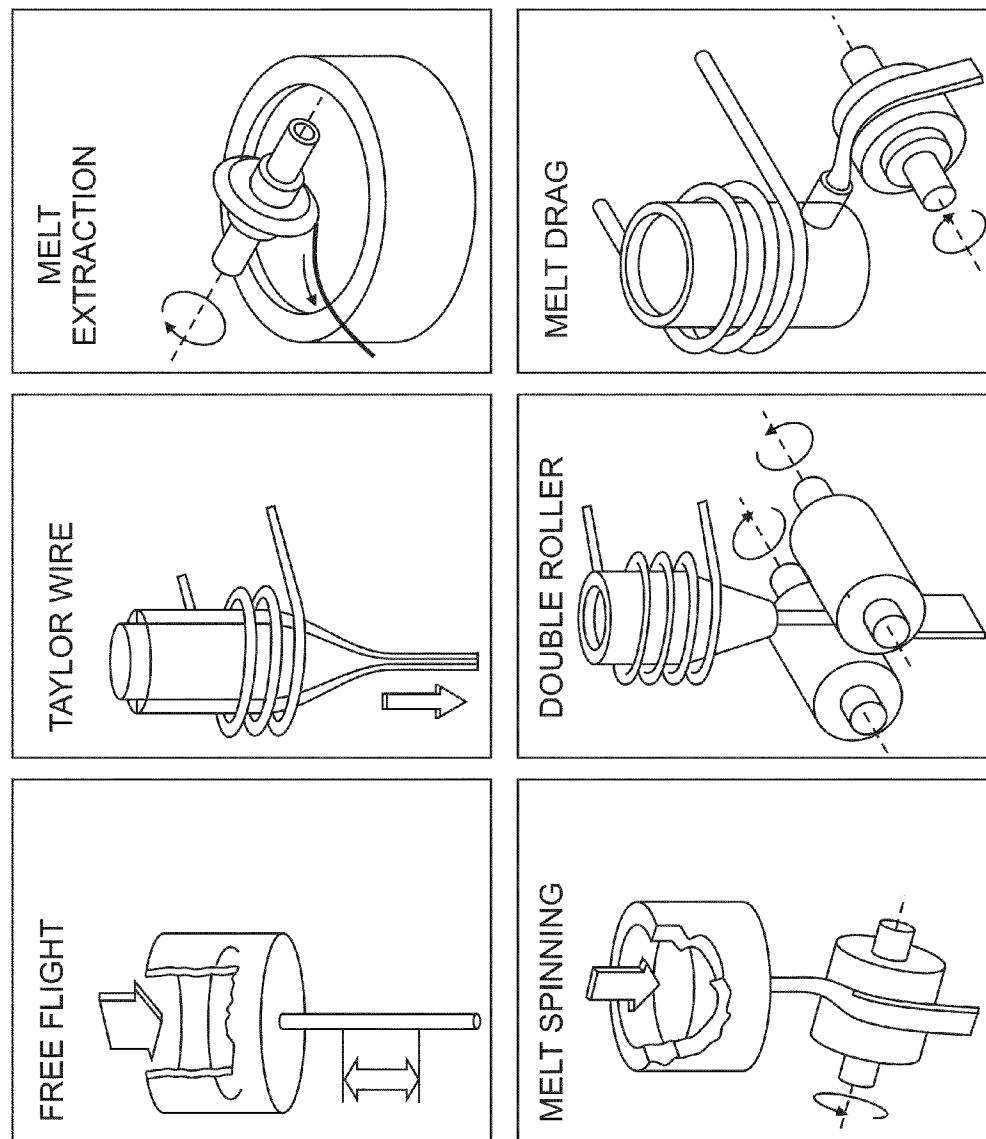


Figure 2e: Melt-Spinning and Melt Extraction



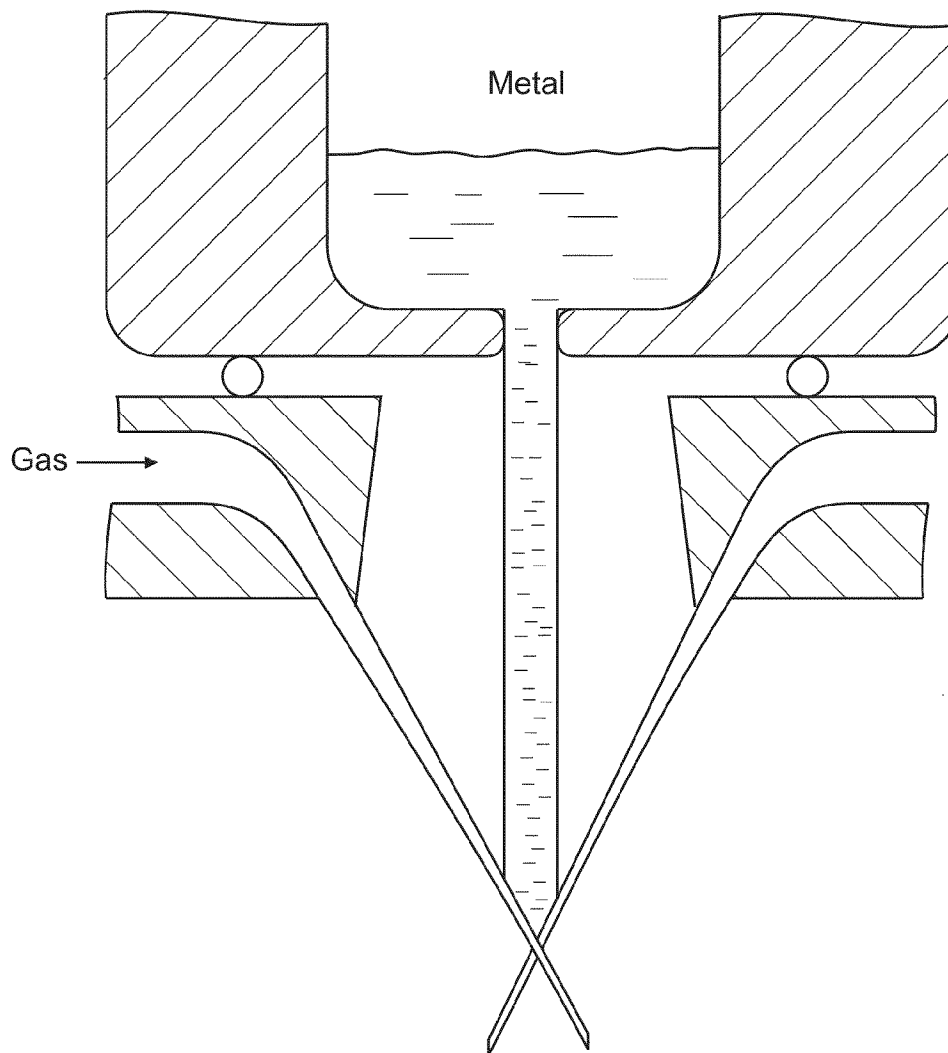


Figure 2f: Gas Atomization

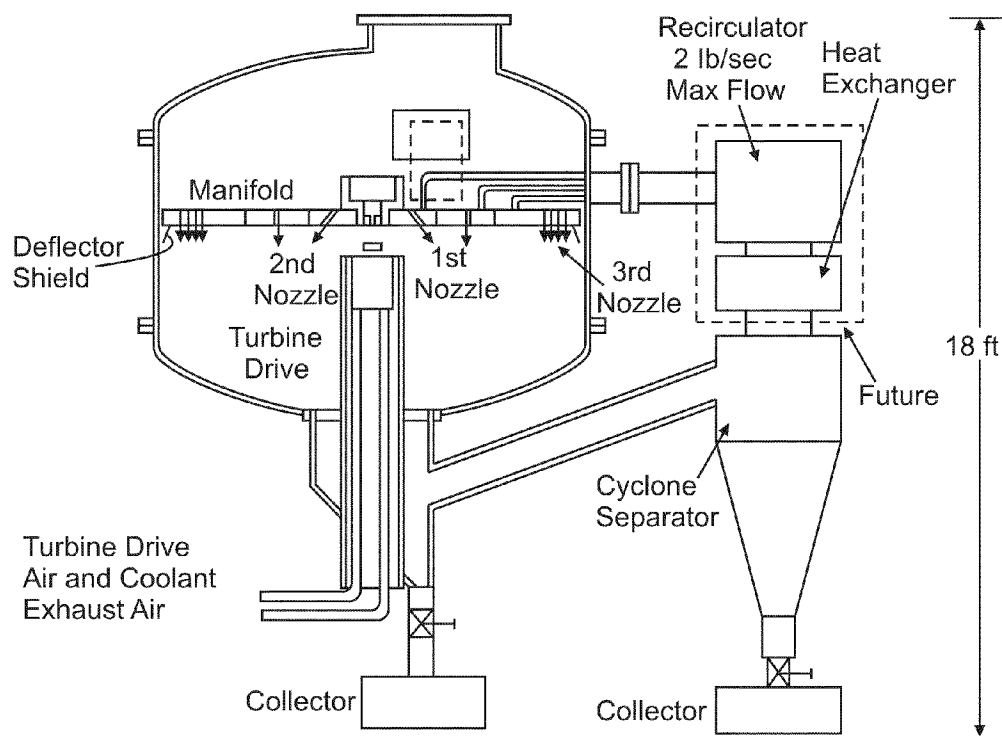


Figure 2g

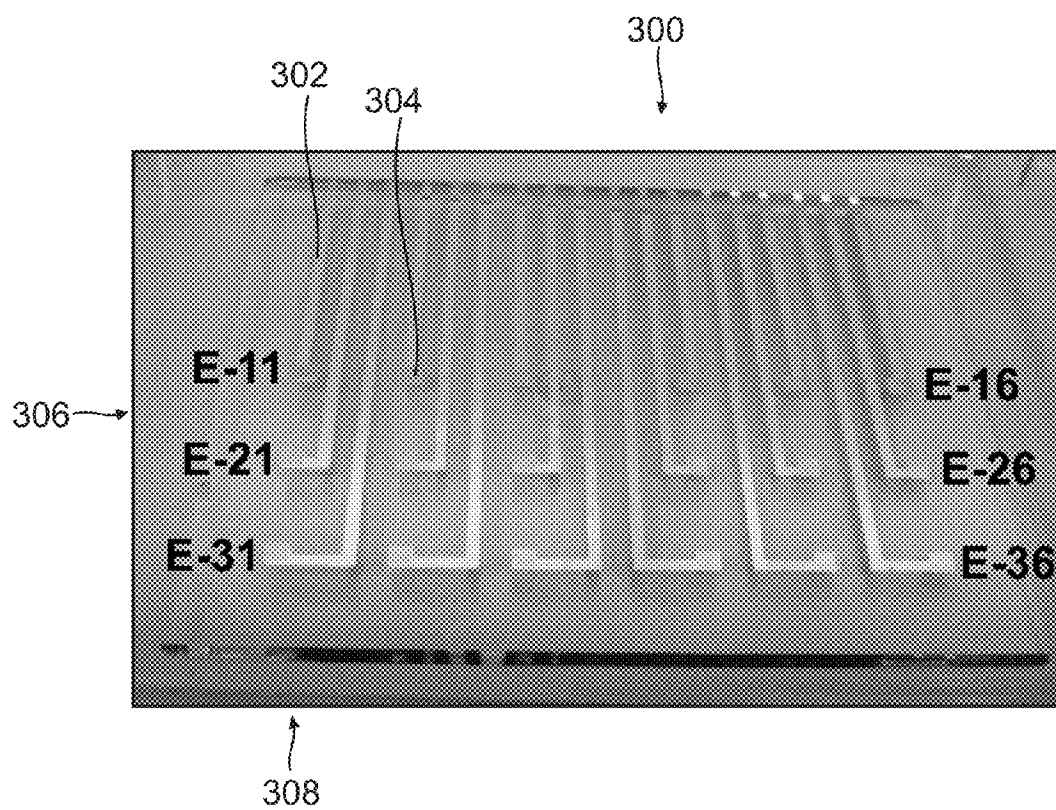


Figure 3

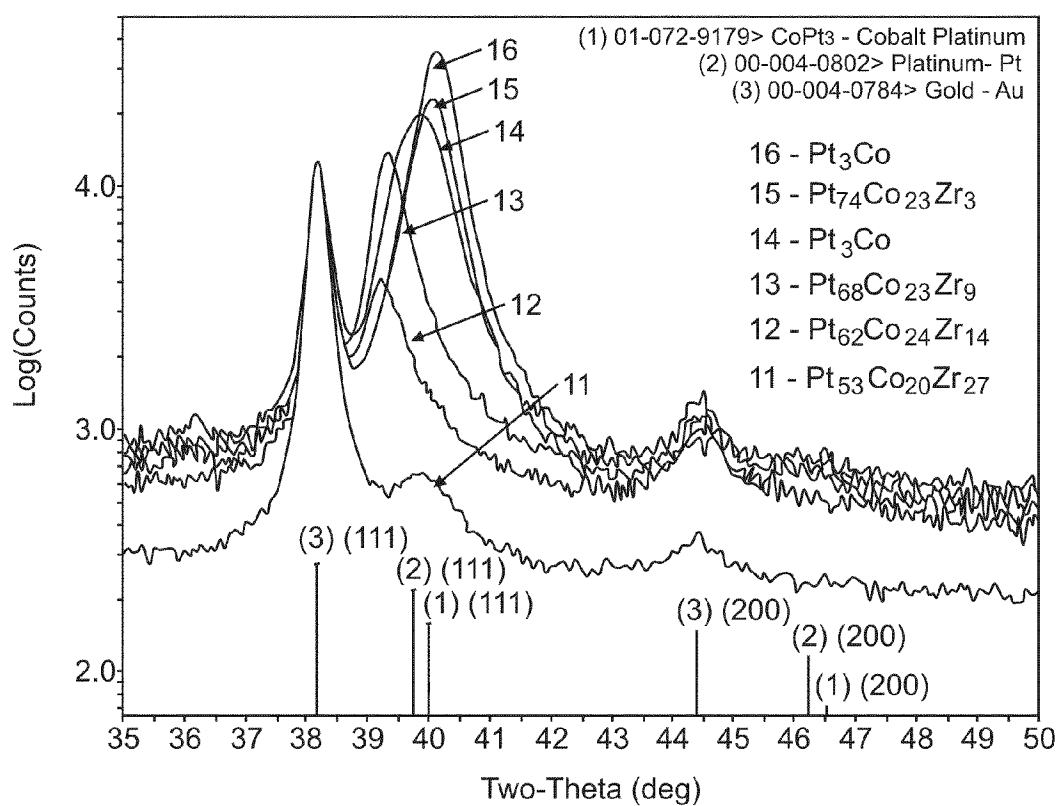


Figure 4a

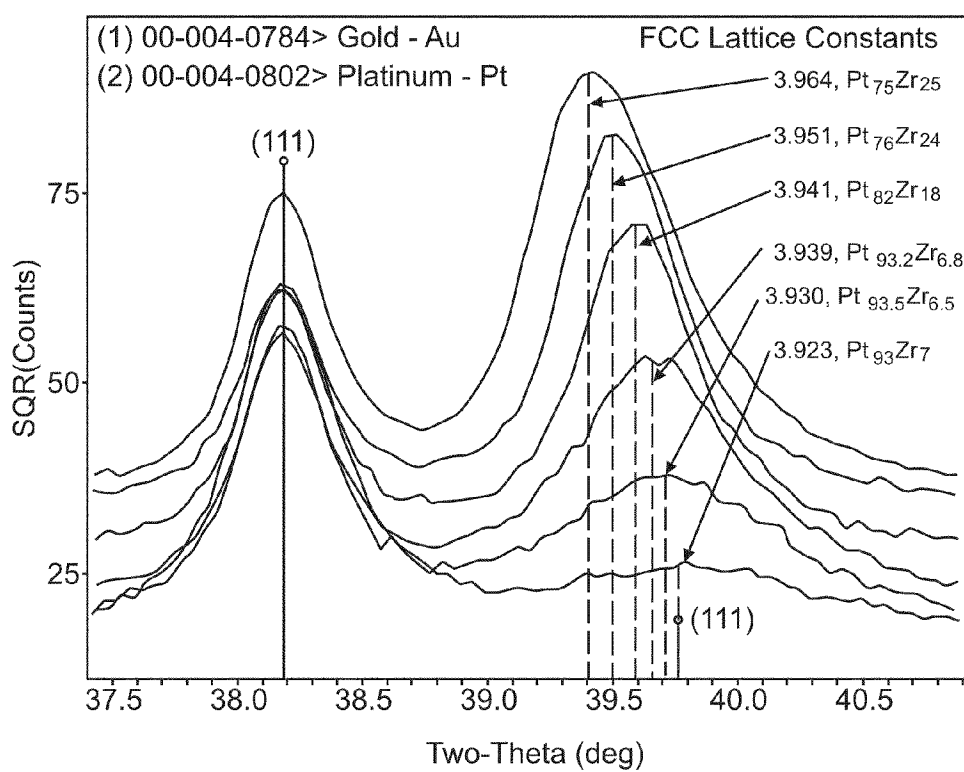


Figure 4b: X-ray Diffraction pattern for Pt<sub>1-x</sub>Zr<sub>x</sub>

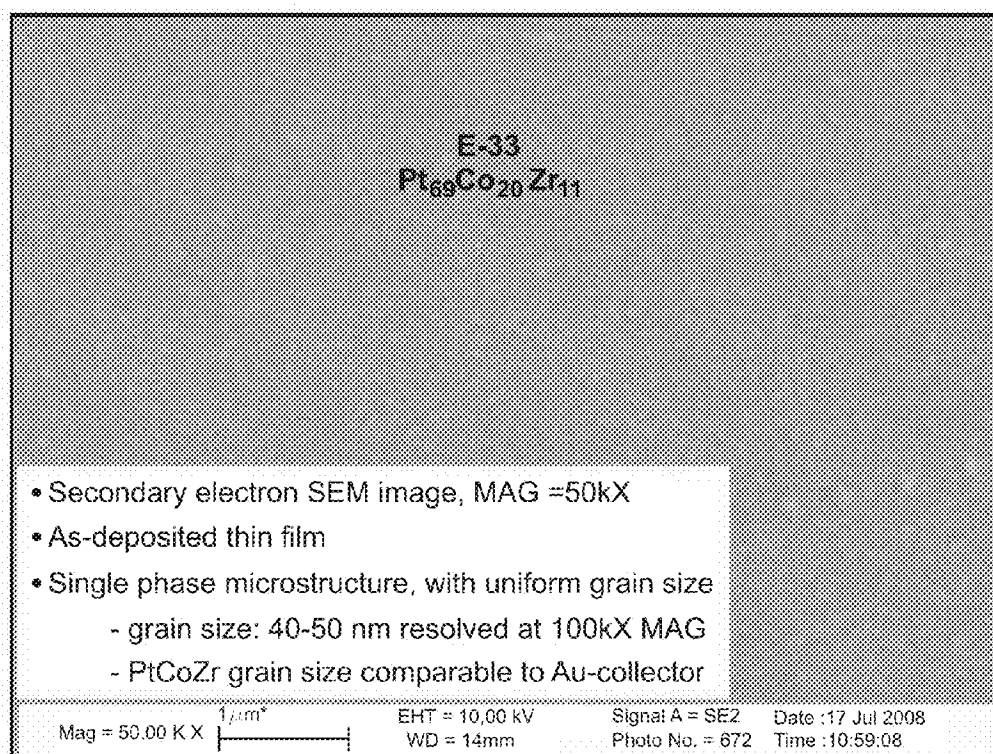


Figure 4c

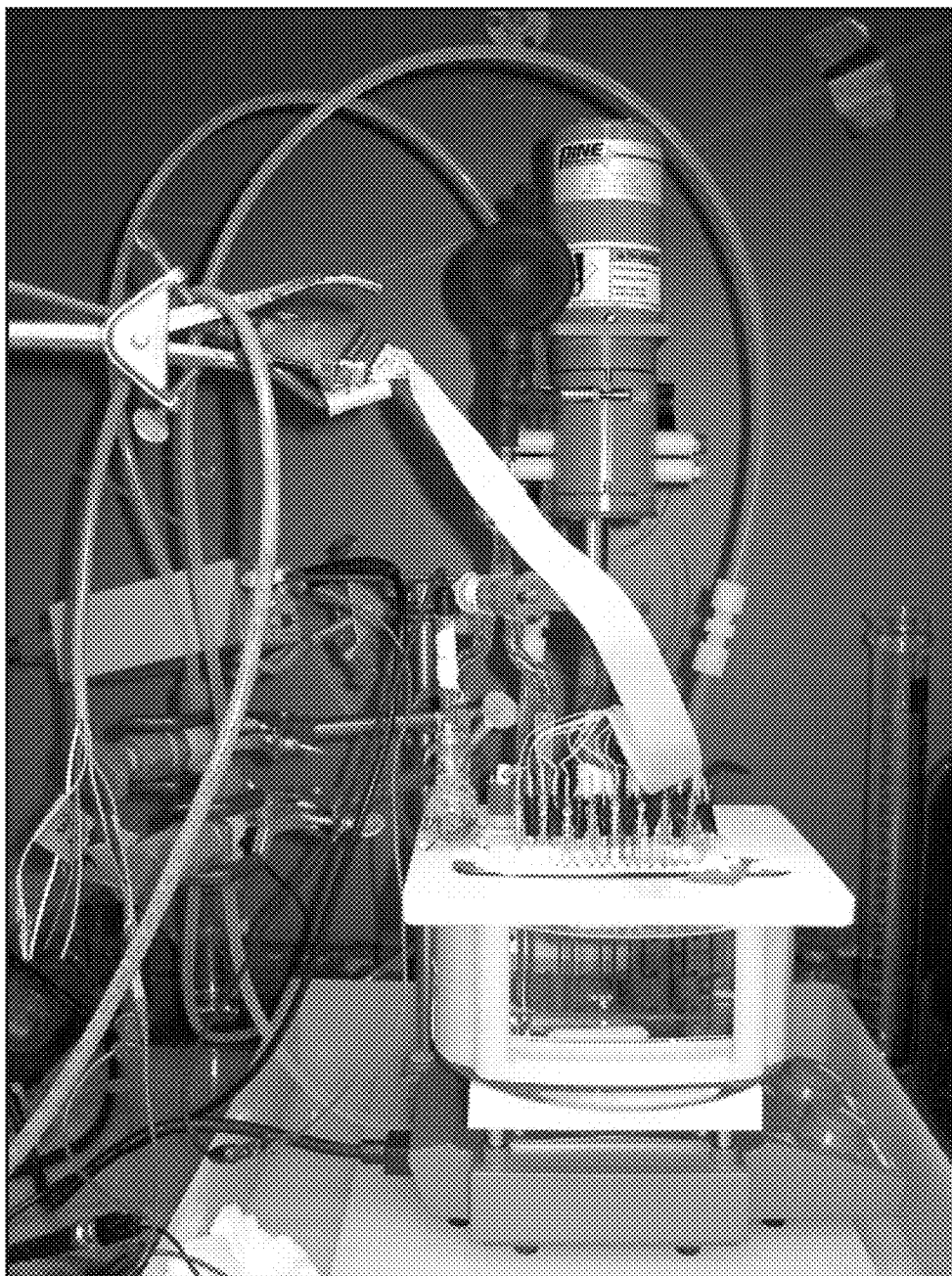


Figure 5a: Custom-built multi-electrode half cell

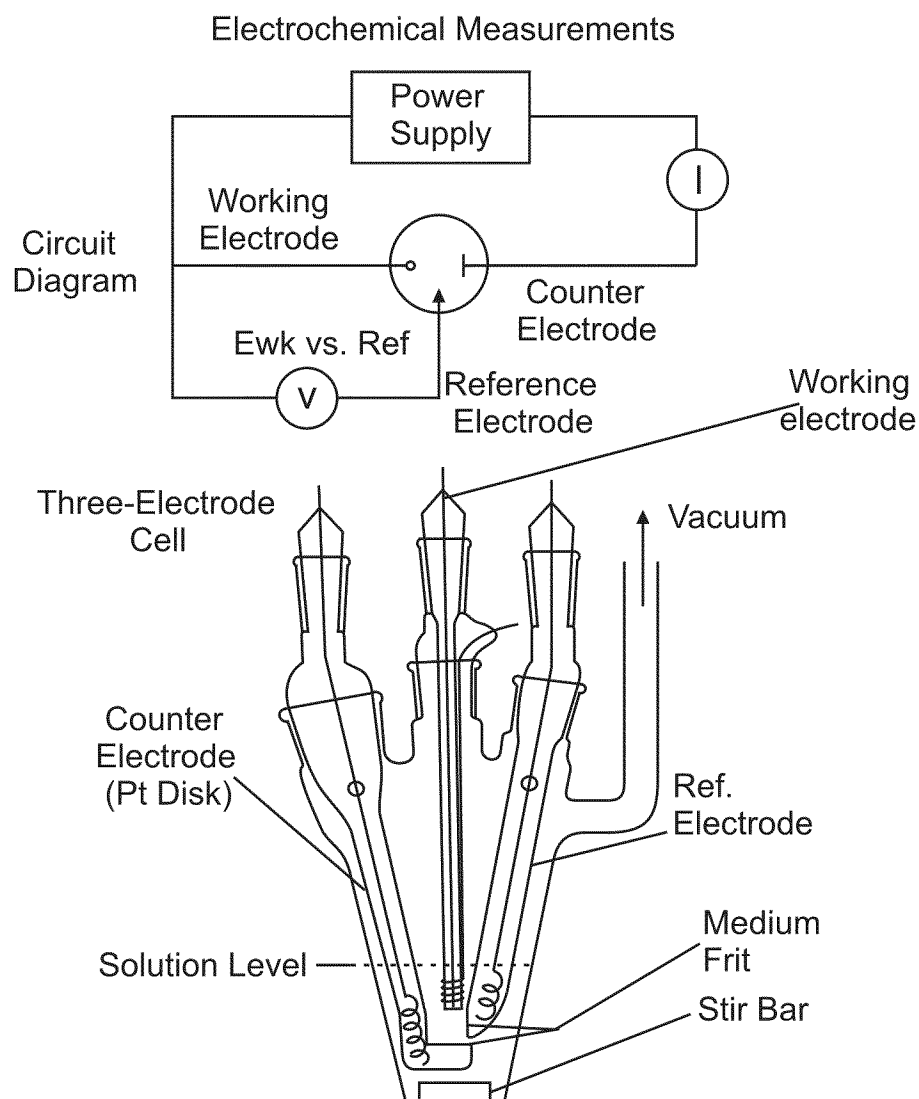


Figure 5b: Schematic of typical half cell



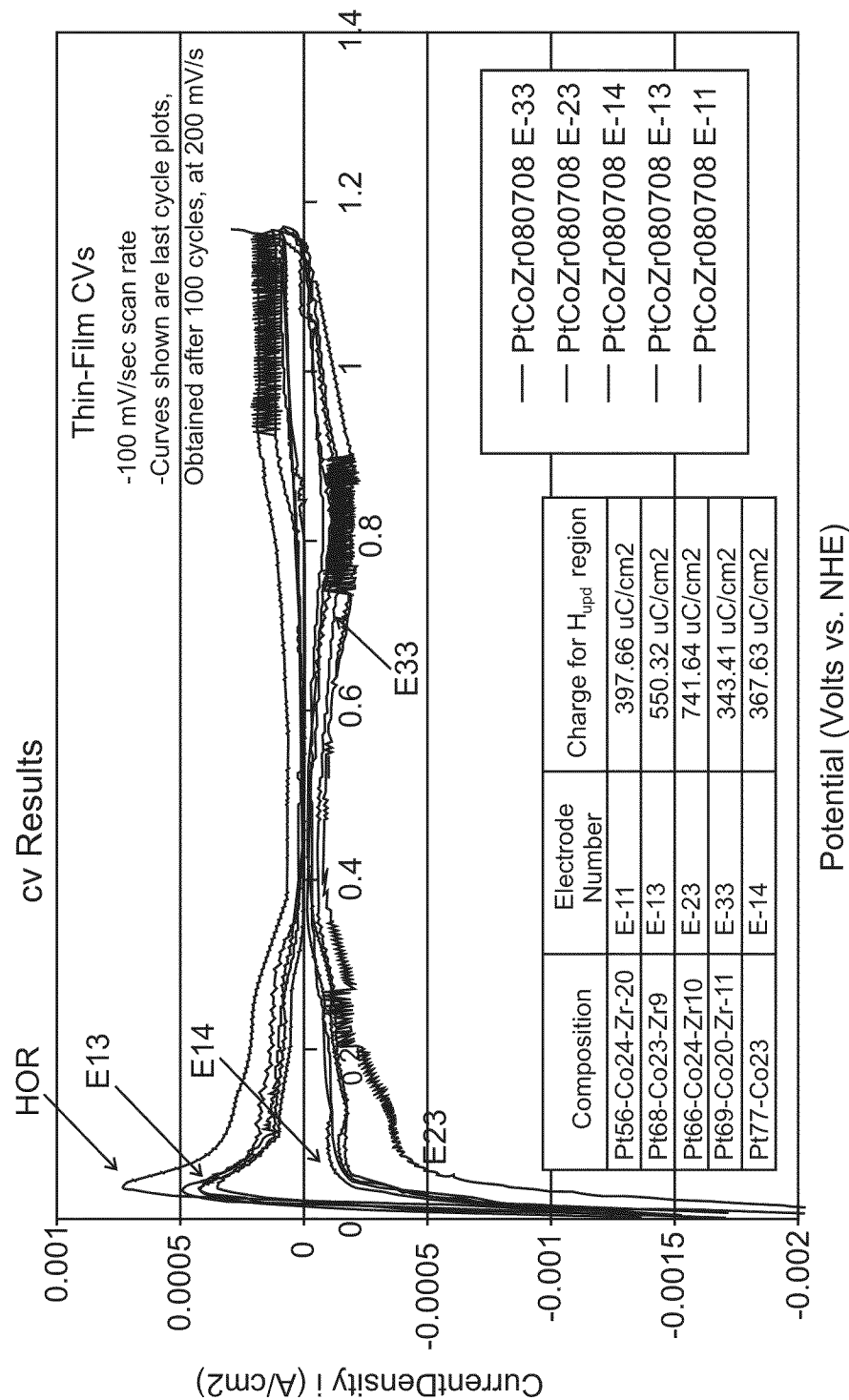


Figure 6

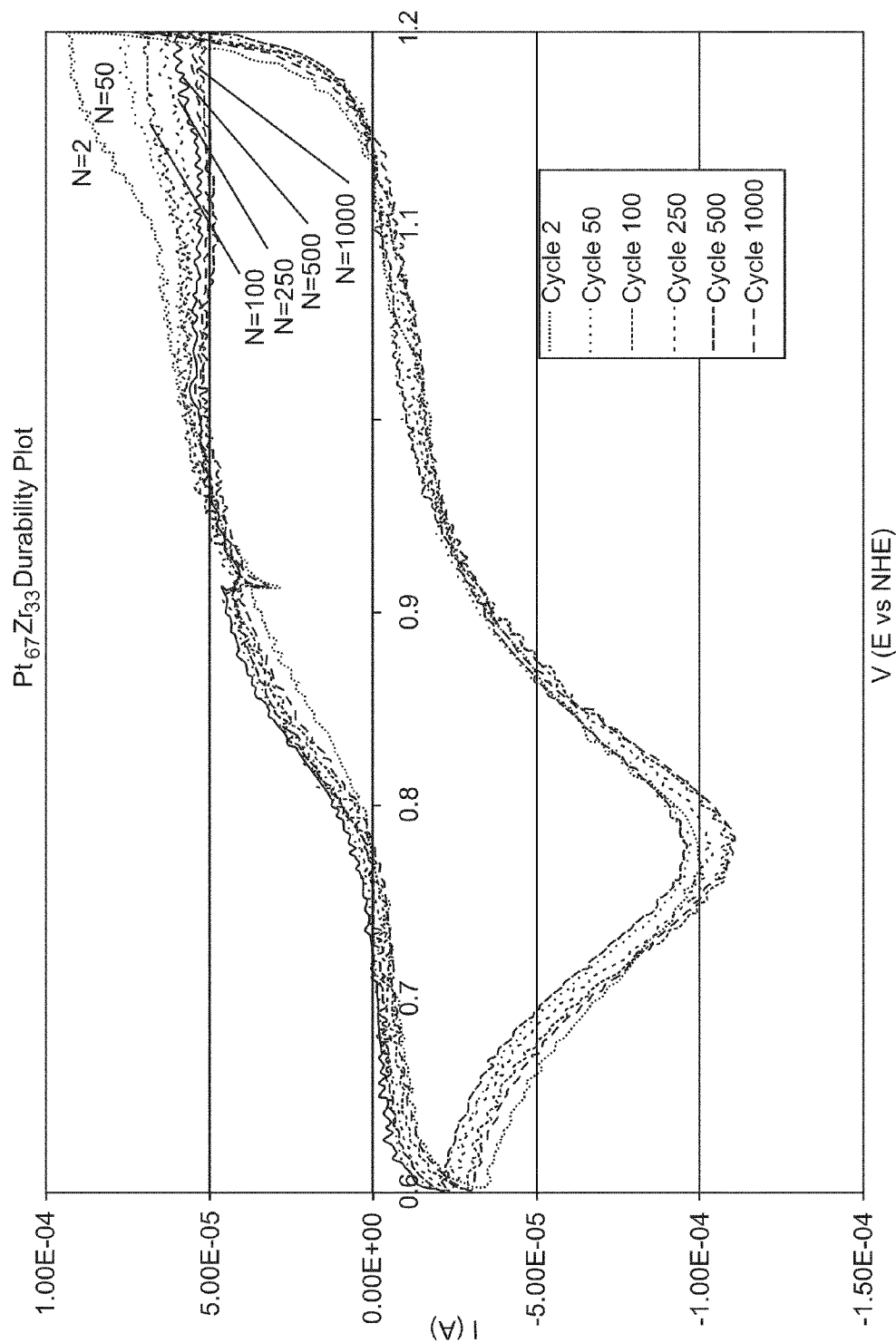
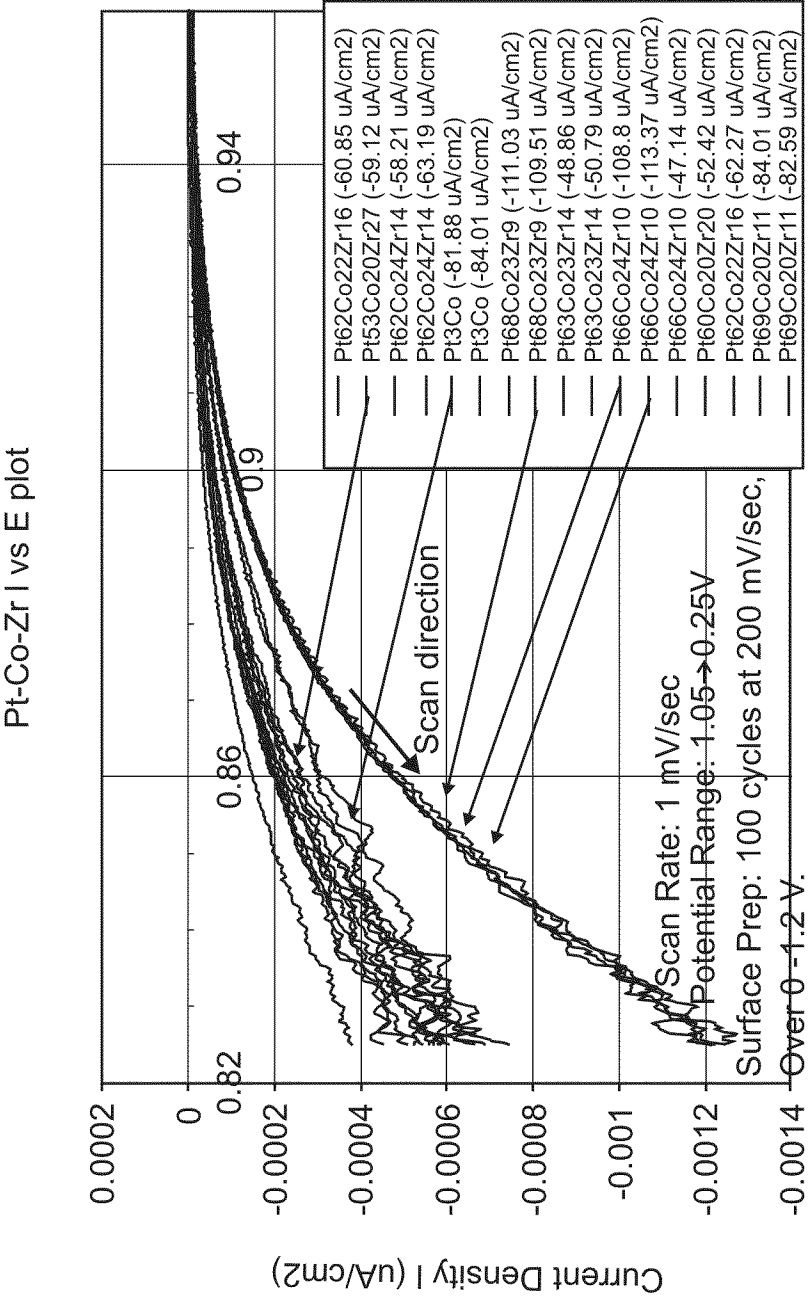
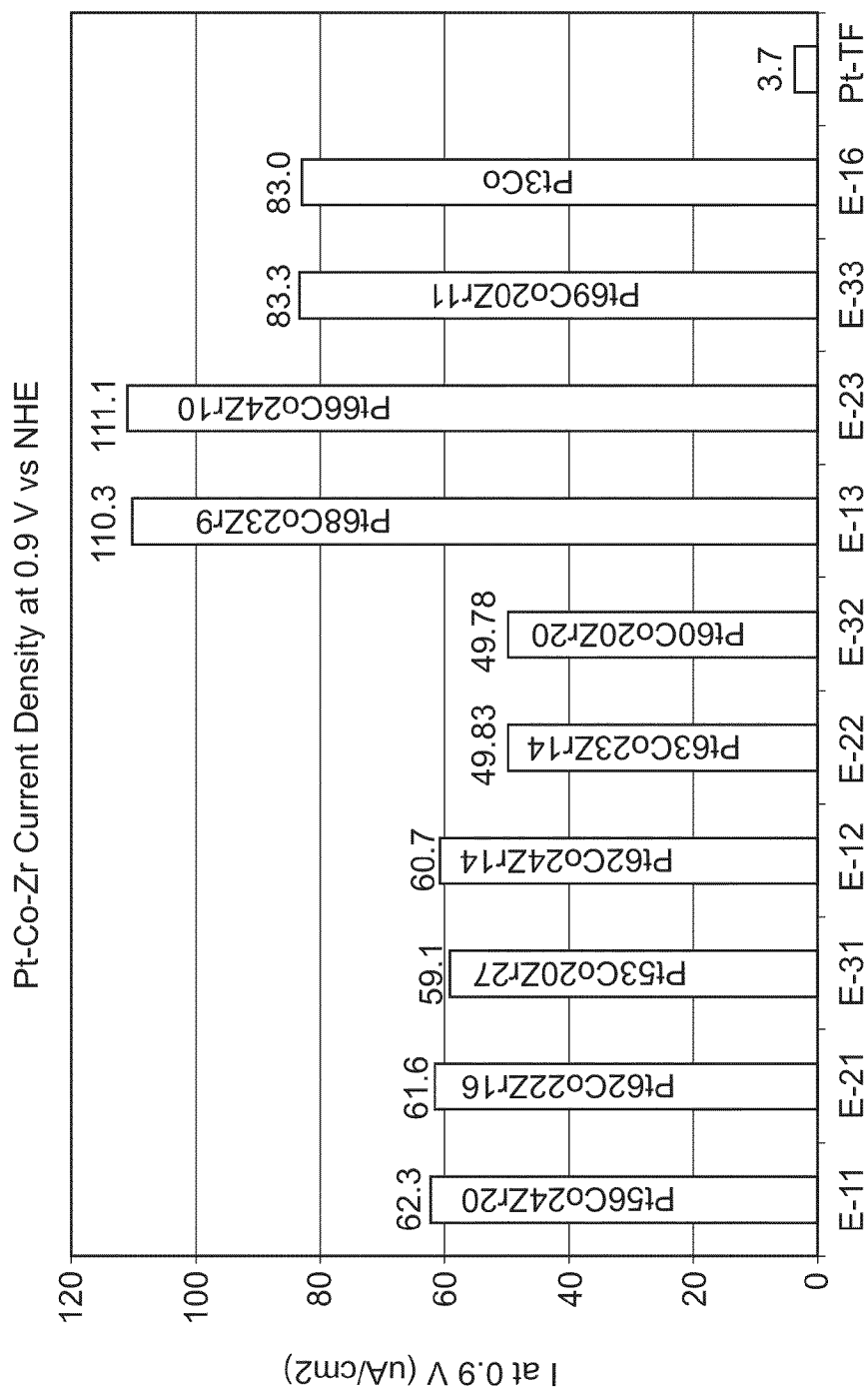


FIG. 7: Durability plot

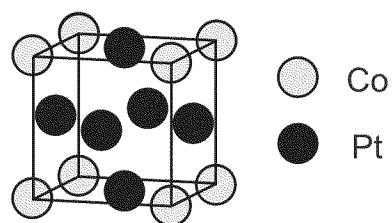
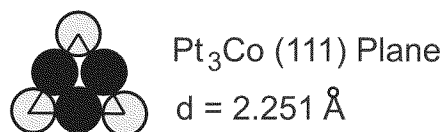


E (V vs NHE)  
ORR Kinetic Current vs. Alloy Composition  
Figure 8



ORR Kinetic Current Density at 0.9 V NHE vs. Alloy Composition

Figure 9

Ordered Pt<sub>3</sub>Co (Cu<sub>3</sub>Au-type)Pt<sub>3</sub>Co (111) Plane

$$d = 2.251 \text{ \AA}$$

At. % Zr	d (Å)	Film Thickness (nm)
0	2.245	190.5
3	2.249	173.5
0 ?	2.261	155.9
9	2.289	109.7
14	2.294	86.8
20	2.262	88.3

Figure 10

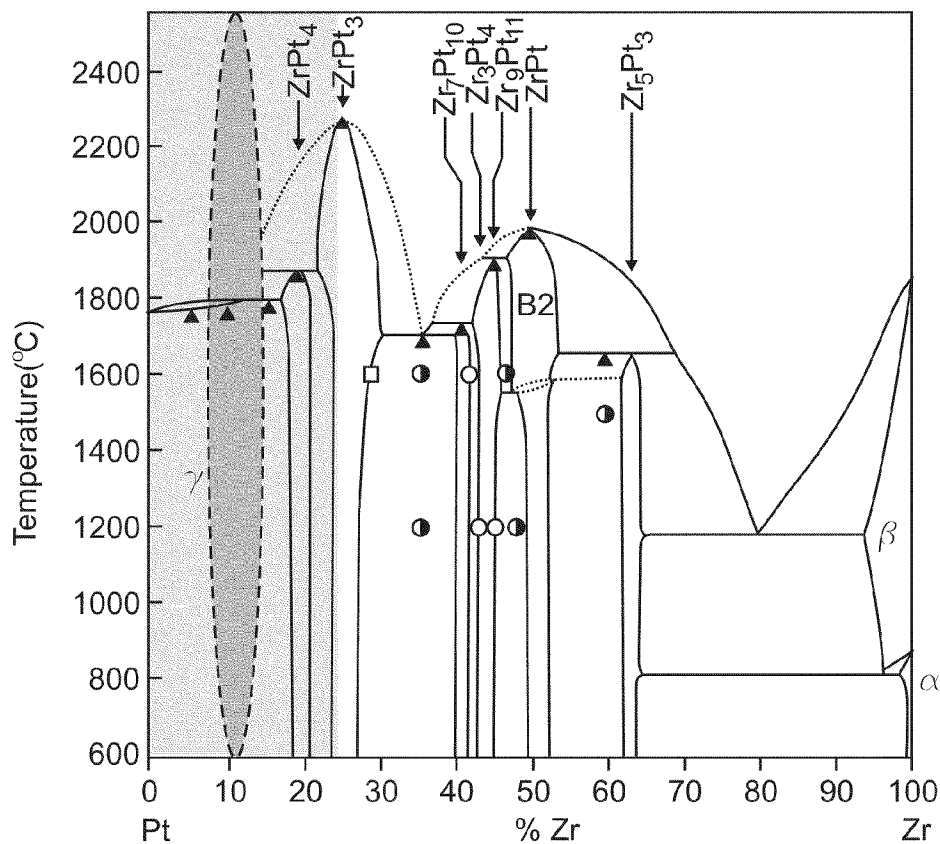
(Pt<sub>3</sub>Co)<sub>1-x</sub>Zr<sub>x</sub> Compositions Mapped on Pt-Zr Binary Diagram

Figure 11a

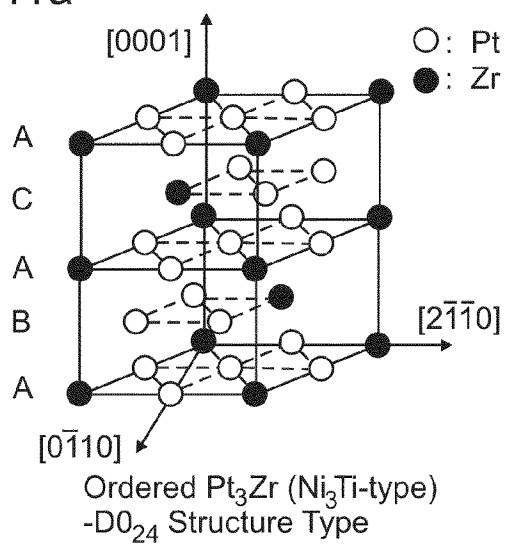


Figure 11b

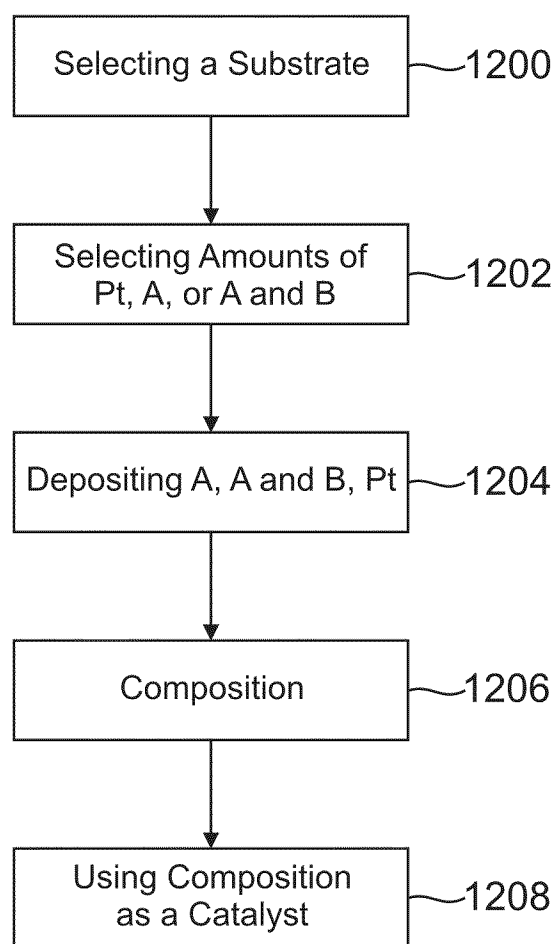


Figure 12

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# NANOSTRUCTURED PLATINUM ALLOYS FOR USE AS CATALYST MATERIALS

## CROSS REFERENCE TO RELATED APPLICATIONS

This application claims the benefit under 35 U.S.C. Section 119(e) of the following co-pending and commonly assigned patent applications which are incorporated by reference herein:

U.S. Provisional Patent Application Ser. No. 61/222,429, filed on Jul. 1, 2009, by Sri R. Narayan and Charles C. Hays, entitled "NANOSTRUCTURED PLATINUM ALLOYS FOR USE AS CATALYST MATERIALS IN FUEL CELLS,"; and

U.S. Provisional Patent Application Ser. No. 61/346,428, filed on May 19, 2010, by Charles C. Hays and Sri R. Narayan, entitled "NANOSTRUCTURED PLATINUM ALLOYS FOR USE AS CATALYST MATERIALS,".

## STATEMENT REGARDING FEDERALLY SPONSORED RESEARCH AND DEVELOPMENT

This invention described herein was made in the performance of work under NASA contract No. NAS7-1407, and is subject to the provisions of Public Law 96-517 (35 U.S.C. 202) in which the Contractor has elected to retain title.

## BACKGROUND OF THE INVENTION

### 1. Field of the Invention

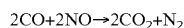
This invention relates to a new composition for use as a catalyst or electro-catalyst material.

### 2. Description of the Related Art

(Note: This application references a number of different publications as indicated throughout the specification by one or more reference numbers within brackets, e.g., [x]. A list of these different publications ordered according to these reference numbers can be found below in the section entitled "References." Each of these publications is incorporated by reference herein.)

Platinum metal, as a discreet nanoparticle or as a film on a nanoparticle support, is the dominant catalyst material for a wide range of catalytic reactions under extreme conditions; e.g., at high temperatures or under acidic environments. Consider the use of Pt catalysts in the Platforming process, first developed in 1949, which enabled the synthesis of gasoline without the addition of lead to the gasoline. Efforts to make the synthesis of gasoline more green include the development of Pt or PtRe catalysts with higher activity or lower Platinum group metal (PGM) loading. In the modern refinery, the Pt-containing catalysts are regenerated once in a six to twenty four month period. The catalyst can be regenerated perhaps 3 or 4 times before it must be returned to the manufacturer for recycling of the PGM catalyst.

In order to reduce vehicle emissions, better catalysts are needed for catalytic converters. In the state-of-the-art (SOA) 3-way catalytic converter, the Pt-based catalysts simultaneously break down the harmful byproducts present in the hot exhaust streams of automobiles exhausts, including: 1) reduction of nitrous oxides to nitrogen and oxygen; 2) oxidation of carbon monoxide to carbon dioxide; and 3) oxidation of unburned hydrocarbon (HCs), according to the following reaction:



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In a 3-way catalytic converter, the Pt or Pt-alloy particles are placed on a high surface area support that is maintained at high temperatures and Pt-based alloys with greater micro-structural stability at these temperatures will improve the converter performance over time.

Fuel cells are widely regarded as an alternative to internal combustion engines, and will play a dominant role in a hydrogen economy as power sources for portable power, transportation, and stationary power applications. However, to meet the future requirements outlined by the US Department of Energy, a new class of catalytic materials is required to improve the performance of electrodes used in advanced fuel cell applications [1]. As such, extensive government and industrial research has been performed in an attempt to commercialize fuel cells. In SOA polymer electrolyte membrane fuel cells (PEMFCs) using an acid polymer electrolyte, platinum (Pt) and platinum group metal (PGM) alloy catalysts are used as the cathode material for the reduction of oxygen, and as the anode material for the oxidation of the hydrogen gas fuel.

In either application, the high cost of Pt is an impediment to their use. A significant amount of research is under way to reduce the Platinum group metal (PGM) content in catalytic converters and fuel cells.

Some challenges limiting the widespread application of PEMFCs, that utilize PGM catalysts are: 1) slow kinetics for oxygen reduction; 2) long-term durability issues manifest by metallurgical effects (e.g., Ostwald particle ripening, and surface area loss due to corrosion); and 3) the high cost of platinum.

The reduced PEMFC durability observed in SOA fuel cell systems is driven in large part by the metallurgical changes in the Pt metal used as the cathode for the oxygen reduction reaction (ORR). During fuel cell operation; grain growth (i.e., Ostwald particle coarsening), corrosion of Pt crystals, and the corrosion/gasification of carbon supports under electrochemical polarization are observed, which collectively result in severe cathode degradation. In addition, the optimal performance of SOA PEMFCs is limited by the sluggish kinetics of the ORR on Pt and its alloys; e.g., a large activation potential (i.e., an over-potential  $V_{over} \sim 300$  mV) exists even for the SOA catalysts. In new the alloyed catalysts, such as Pt<sub>3</sub>Co, the durability of polymer electrolyte cell membrane is reduced due to poor corrosion resistance of the cell electrodes, where Co is corroded from the catalyst surface (cathode), and ultimately crosses over into the membrane [2].

In this application, embodiments of the present invention teach a new class of Pt-based catalyst materials. In the description of the catalyst performance, focus on the use of these materials as the cathode material in a PEMFC, and the results presented show that their performance far exceeds that demonstrated by SOA Pt materials, with a greatly reduced Pt content or Pt loading. However, other applications are also disclosed.

## SUMMARY OF THE INVENTION

There is provided in the practice of embodiments of this invention, a series of binary and ternary Pt-alloys, that promote the important reactions for catalysis at an alloy surface; oxygen reduction, hydrogen oxidation, and hydrogen and oxygen evolution. The first two of these reactions are essential when applying the alloy for use in a PEMFC.

To overcome the limitations in the prior art described above, and to overcome other limitations that will become apparent upon reading and understanding the present invention, the present invention discloses a platinum (Pt) metal



containing alloy composition useful as a catalyst electrode, comprising a film comprising a compound of Pt and at least one early transition metal element A from group IVB or VB of the Periodic table.

The film may comprise a continuous film on a nanoparticle support, a continuous film on a hollow spherical nanoparticle (e.g., nanoshell) support, a continuous film on a micron-sized metallic or non-metallic support, or a continuous film on a wire or wire-gauze support, for example.

There may be less than 50% Atomic % (At. %) Pt in the compound.

Element A may comprise at least one valve metal element of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb). A may comprise at least one of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb), and B may comprise at least one of Cobalt (Co), Nickel (Ni), and Iron (Fe), and with at least one PGM element of Palladium (Pd), Ruthenium (Ru), Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir).

The Pt-containing compound may further comprise at least one late transition metal B, thereby forming a Pt—B—A alloy. For example, A may comprise at least one of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb), and B may comprise at least one of Cobalt (Co), Nickel (Ni), and Iron (Fe). The alloy may be  $(\text{Pt}_3\text{Co})_{100-y}\text{Zr}_y$ , with  $0 \leq y \leq 30$  (At. %). The alloy may comprise  $(\text{Pt}_{100-x}\text{Co}_x)_{100-y}\text{Zr}_y$ , with  $0 \leq x \leq 80$  and  $0.5 \leq y \leq 60$ .

The Pt-containing compound may further comprise at least one platinum group metal (PGM), thereby forming a Pt—PGM—B—A alloy, wherein the PGM element comprises at least one of Palladium (Pd), Ruthenium (Ru), Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir).

The alloy may be crystalline, for example, nanocrystalline with a grain size of no more than 100 nm, or nanocrystalline with a preferred grain size of less than 10 nm. The alloy may have a (111) crystallographic orientation.

An amount of A and an amount of Pt in the film may be such that the film is at least 2 times more electrochemically active, in an oxygen reduction reaction (ORR), than Pt. An amount of A and the amount of Pt in the film may be such that the film is electrochemically stable, with no decrease in electrochemical activity of the film used with an electrolyte. To determine the durability of said alloy, the electrochemical activity is characterized by conducting multiple cyclic voltammograms (CV) over a potential range of 0.0 to at least 1.2 Volts (vs. NHE), with a potential scan rate of the order 100 mV/sec, and a total number of cycles exceeding  $N=1000$ . For durable compositions, the current density at 0.9 V (vs. NHE) should not decrease on electrochemical cycling, within the measurement accuracy (e.g.,  $\pm 10$  microamps).

The electrolyte may be an acid electrolyte. The acid electrolyte may be a mixture of perchloric acid in water ( $\text{HClO}_4/\text{H}_2\text{O}$ ) (e.g., a 1 molar perchloric acid concentration in water ( $\text{HClO}_4/\text{H}_2\text{O}$ )), or mixture of sulfuric acid in water ( $\text{H}_2\text{SO}_4/\text{H}_2\text{O}$ ) (e.g., a 1 molar sulfuric acid concentration in water ( $\text{H}_2\text{SO}_4/\text{H}_2\text{O}$ )). The acid electrolyte may saturate a polymer exchange membrane electrolyte (e.g. Nafion).

For example, synthesized Pt—Co—Zr thin films are stable in 1 M perchloric acid and are electrochemically active for the oxygen reduction reaction, with kinetic currents at 0.9 V (that greatly exceed those of Pt, by amounts as great as thirty times [30x]).

The present invention further discloses a platinum (Pt) metal containing alloy composition useful as a catalyst electrode, comprising a nanoparticle comprising a compound of Pt and at least one early transition metal element A from group IVB or VB of the Periodic table.

The present invention further discloses a platinum (Pt) metal containing alloy composition useful as a catalyst electrode, comprising a metallic nanocrystalline cluster or quantum dot on a nanoparticle support, wherein the metallic nanocrystalline cluster or quantum dot comprises a compound of Pt and at least one early transition metal element A from group IVB or VB of the Periodic table.

#### BRIEF DESCRIPTION OF THE DRAWINGS

Referring now to the drawings in which like reference numbers represent corresponding parts throughout:

FIG. 1 is a fuel cell comprising electrodes according to the present invention;

FIGS. 2a-2g are schematics of co-sputtering, electron beam evaporation, vapor phase condensation, atomic layer deposition, melt spinning and melt extraction, gas atomization, and rotary atomization processes, that may be used to fabricate the compositions of the present invention;

FIG. 3 is a photograph of an 18 electrode Pt—Co—Zr thin film array, showing electrodes upon which samples E11-E36 are deposited;

FIGS. 4a and 4b show X-ray diffraction results for  $\text{Pt}_3\text{Co}$  to  $\text{Pt}_{53}\text{Co}_{20}\text{Zr}_{27}$  films, and  $\text{Pt}_{100-x}\text{Zr}_x$  films, respectively, and FIG. 4c is a secondary electron scanning electron microscope (SEM) image of the top surface of sample E33, a  $\text{Pt}_{69}\text{Co}_{20}\text{Zr}_{11}$  as deposited thin film (50,000x magnification), showing single phase microstructure, with uniform grain size (40-50 nm resolved at 100,000x magnification), wherein the PtCoZr film's grain size and crystallographic orientation (111) are comparable to the grain size of the underlying Au-film on which it was deposited;

FIG. 5a is an image of an electrochemical measurement set up and FIG. 5b is a schematic of a typical electrochemical half cell;

FIG. 6 shows voltammograms (CV), plotting current (amps) as a function of voltage applied to PtCoZr and PtCo films (voltage with respect to a normal hydrogen electrode defining 0 volts), for samples E11, E13, E23, E33 and E14 using the PtCoZr multi-electrodes array of FIG. 3 and the multi electrode half cell set up of FIG. 5a, wherein the voltage is scanned at a 100 mV/second scan rate and the curves shown are the last cycle plots obtained after 100 cycles; the surfaces were prepared by conducting 100 cycles at 200 mV/second scan rate over the same potential range;

FIG. 7 is a durability plot for a  $\text{Pt}_{67}\text{Zr}_{33}$  film, showing multiple cyclic voltammograms taken after the film has been electrochemically prepared (by e.g., by cycling voltage over 100 cycles), with the individual voltammograms shown after conducting  $N=2, 50, 100, 250, 500$ , and 1000 cycles;

FIG. 8 shows nearly potentiostatic polarization curves in the kinetic region for the ORR, for various PtCoZr and PtCo compositions, wherein current density (microamps/cm<sup>2</sup>) is plotted as a function of voltage applied to the PtCoZr and PtCo films (voltage with respect to a normal hydrogen electrode defining 0 volts), the voltage is ramped from 1.05 V to 0.25 V at a 1 mV/second scan rate, and the measurements of FIG. 8 are taken after the surface of the films has been prepared by performing 100 voltage cycles over 0 V to 1.2 V at a 200 mV/second scan rate;

FIG. 9 shows ORR current density (at 0.9 V vs. NHE) for different alloy compositions of the present invention, wherein the ORR current density for a Pt thin film (Pt-TF) is also shown;

FIG. 10 is a schematic of the binary  $\text{Pt}_3\text{Co}$  composition showing the 111 plane;

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FIG. 11a is a graph showing the binary phase diagram for the Pt—Zr alloy series, at which various  $\text{Pt}_{100-x}\text{Zr}_x$  compositions occur, wherein the ellipse shows compositions with high electrochemical performance, and FIG. 11b is a schematic showing the  $\text{D0}_{24}$  crystal structure for the  $\text{Pt}_3\text{Zr}$  composition; and

FIG. 12 is a flowchart illustrating a method of fabricating and using a Pt metal alloy composition as a catalyst.

## DETAILED DESCRIPTION OF THE INVENTION

In the following description of the preferred embodiment, reference is made to the accompanying drawings which form a part hereof, and in which is shown by way of illustration a specific embodiment in which the invention may be practiced. It is to be understood that other embodiments may be utilized and structural changes may be made without departing from the scope of the present invention.

### Technical Description

Two series of Pt-based alloys, in thin film form, have been prepared by co-sputtering. As examples, the present invention shows that members of two platinum-metal containing composition manifolds are stable in acid solutions, and exhibit electro-catalytic performance that meets or exceeds that of pure Pt metal. Examples of the two platinum metal containing composition manifolds are A)  $(\text{Pt}_{100-x}\text{Co}_x)_{100-y}\text{Zr}_y$ ,  $0 \leq x \leq 80$ , and  $0.5 \leq y \leq 60$  (At. %); and B)  $\text{Pt}_{100-x}\text{Zr}_x$ ,  $8 < x < 33$  (At. %).

Thus, the compositions of the present invention may be useful in any applications that would benefit from electro-catalytic performance. For example the new Pt-based alloy catalysts of the present invention may be useful as electrodes in a fuel cell, as illustrated in FIG. 1. The fuel cell typically comprises an electrolyte between the anode and the cathode. The composition of the present invention may be used as a cathode and/or an anode in the fuel cell.

### Fabrication

Combinatorial film deposition methods, as described in [1], were used to simultaneously prepare a wide range of Pt—Co—Zr compositions for evaluation. Using a high-throughput, co-sputtering, synthesis technique, an array of thin film specimens in the ternary alloy series  $(\text{Pt}_3\text{Co})_{100-x}\text{Zr}_x$ ,  $0 \leq x \leq 30$  (At. %), were simultaneously prepared.

FIG. 2a shows a schematic of the co-sputtering process. All films were co-sputtered from two targets, each made from research grade materials, with minimum purities of 99.99%;  $\text{Pt}_3\text{Co}$  and Zr (Kurt J. Lesker). A typical co-sputtering procedure comprises evacuating the sputtering chamber to a base chamber pressure of  $<1 \times 10^{-6}$  Torr, followed by film deposition carried out under an Argon gas pressure of 15 mTorr.

A multi-electrode array comprising of 18 thin film electrodes were deposited using a three-step physical masking method. First, an 18-segment current collector array was fabricated using a nanostructured Au thin film over a Ti adhesion layer. The average Au grain size was 40-50 nm. The Au/Ti films were sputtered through a patterned mask onto a polished borosilicate glass substrate, with nominal dimensions of  $4'' \times 5'' \times 1/8''$ . The Au films were strongly oriented, with a (111) crystallographic orientation.

The Pt—Co—Zr catalyst films were co-sputtered after physically masking off an equally spaced array of  $1/8'' \times 1/8''$  openings, each placed above the Au current collector structure. The individual films were deposited onto an 18-segment current collector structure comprised of nanostructured Au thin films, with average Au grain size of 40-50 nm.

FIG. 2b illustrates an electron beam evaporation apparatus [12] that can be used to deposit the Pt films of the present

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invention, comprising a heatable substrate holder, quartz crystal, shutter, crucible, and 4-pocket e-beam evaporator.

FIG. 2c illustrates a vapor phase condensation apparatus [6] that can be used to synthesize the Pt alloys of the present invention in particulate form, comprising a vacuum chamber with computerized process control comprising a heating element ( $e^-$  beam evaporator), power supply, powder, quenchant gas. Powder collection and final packaging is also shown.

FIG. 2d illustrates an atomic layer deposition (ALD) apparatus [7] (fluidized bed ALD reactor schematic) that can be used to deposit the Pt films of the present invention, comprising Mass Flow Controllers (MFC), Nitrogen Source  $\text{N}_2$ , reagent reservoirs A and B, vibrational motors, fluidized bed reactor, and vacuum pump.

FIG. 2e illustrates an melt spinning and melt extraction apparatus [8] that can be used to synthesize the Pt alloys of the present invention in thin ribbon or fine wire form, comprising melt spinning, double roller, melt drag, free flight, taylor wire, and melt extraction.

FIG. 2f illustrates Gas Atomization apparatus [9] that can be used to synthesize the Pt alloys of the present invention in powder form, comprising gas and metal source.

FIG. 2g illustrates a centrifugal atomizer [10] that can be used to synthesize the Pt alloys of the present invention in powder form, comprising turbine drive air and coolant exhaust air, collectors, deflector shield, manifold,  $2^{nd}$  nozzle,  $1^{st}$  nozzle, recirculator, heat exchanger, and cyclone separator.

FIG. 3 shows the collector structure 300 comprising 18 thin film electrodes 302 of Au thin films disposed in an array, with different PtCo and PtCoZr samples 304 deposited thereon. The samples 304, labeled E-ab depending on their position on the array, comprise different compositions having different amounts of Pt, Co and Zr, where “a” corresponds to the row number 306 and “b” corresponds to the column number 308. For example, sample E-11 is positioned in the first row and first column.

### Experimental Characterization of the Films Structure

FIGS. 4a and 4b illustrate X-ray diffraction (XRD) data for a platinum (Pt) metal containing alloy composition for use as a catalyst electrode, comprising a film including a compound of Pt and at least one early transition metal element A from group IVB or VB of the Periodic table. In this example, the Pt-containing compound further comprises at least one late transition metal B, thereby forming a Pt—B-A alloy.

FIG. 4a shows the Au films were strongly oriented, with a (111) crystallographic orientation. Also shown in FIG. 4a are the XRD patterns for six films from the array, wherein nominal compositions from  $\text{Pt}_3\text{Co}$  (16) to  $\text{Pt}_{53}\text{Co}_{20}\text{Zr}_{27}$  (11 or E-11) are shown. Each of the Pt—Co—Zr thin films exhibits a (111) crystallographic orientation. Some additional lines in the spectra indicate the presence of some grains with differing orientation. The decreased intensity for  $\text{Pt}_{53}\text{Co}_{20}\text{Zr}_{27}$  is consistent with the reduced thickness of the films in this part of the array.

As shown in FIG. 4a, the X-ray spectra reveal that there appears to be a smooth increase in the lattice parameter with increasing x, moving from  $\text{Pt}_3\text{Co}$  ( $a=2.245$  Å,  $\text{Cu}_3\text{Au}$  structure type) to  $\text{Pt}_{62}\text{Co}_{24}\text{Zr}_{14}$  ( $a=2.294$  Å). For  $x>14\%$ , the lattice parameter decreases.

As shown in FIG. 4b, the X-ray spectra for  $\text{Pt}_{100-x}\text{Zr}_x$  reveal that there appears to be a smooth decrease in the lattice parameter with decreasing x, moving from  $\sim\text{Pt}_3\text{Zr}$  ( $a=3.964$  Å,  $\text{Ni}_3\text{Ti}$  structure type) to  $\text{Pt}_{93}\text{Zr}_7$  ( $a=3.923$  Å, Pt-ric,  $\text{Pt}_4\text{Zr}$  structure type [i.e.,  $\text{Cu}_3\text{Au}$  type]).

FIG. 4c shows SEM images of the compositions of the present invention, showing single phase microstructure, with uniform grain size (40-50 nm).

#### Electrochemical Properties

The electrochemical properties of the  $(\text{Pt}_3\text{Co})_{1-x}\text{Zr}_x$  films were measured using a high-throughput, multi-electrode, screening technique developed at Nasa's Jet Propulsion Laboratory. This technique enables the simultaneous evaluation of polarization behavior, active area, and durability for multiple thin film specimens.

FIGS. 5a and 5b illustrate electrochemical measurement set ups used for the measurements of the present invention [11].

The results of these electrochemical measurements show that a wide range of compositions within the  $(\text{Pt}_3\text{Co})_{100-x}\text{Zr}_x$  composition manifold are stable in acid solutions (e.g., 0.1 M  $\text{HClO}_4/\text{H}_2\text{O}$  electrolytes). This stabilization has been achieved by the addition of early transition-metal (ETM) elements from groups IVB and VB of the periodic table (e.g., Ti, Zr, Hf, and Nb). The addition of the ETM element(s) also enables the addition of late transition metal (LTM) moieties (e.g., Ni, Co, and Fe), thereby reducing the Pt-metal content further.

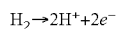
In order to exhibit electro-catalytic behavior, the films of the present invention were electrochemically prepared by applying a voltage to the films (with respect to a normal hydrogen electrode (NHE) defining 0 volts), wherein the voltage was ramped from 0 to 1.2 V and the ramping was repeated or cycled over a sufficient number of cycles (approximately 100 cycles). After sufficient number of cycles (e.g., 100 cycles), the film exhibits electrochemical characteristics.

#### Anodic Behavior

The films of the present invention may be used as an anode to catalyze a hydrogen oxidation reaction.

Electro-catalytic behavior is demonstrated by the results presented in FIG. 6, which shows a cyclic voltammogram (CV), taken using the set ups in FIGS. 5a and 5b, for the compositions E11, E13, E23, E33 and E14 which exhibit a (111) crystallographic orientation. This CV was conducted at a scan rate of 100 mV/sec, in a de-aerated 0.1 M  $\text{HClO}_4/\text{H}_2\text{O}$  electrolyte. Note that the hydrogen oxidation peaks are those associated with the Pt(111) or PtCo(111) crystal faces [active Pt-site area values in Table-1].

The hydrogen oxidation reaction (HOR) peak is a direct measurement of the oxidation of hydrogen to produce electron(s) ( $e^-$ ) and a hydrogen nucleus ( $\text{H}^+$ ), catalyzed by the composition of the present invention acting as an electrode (anode), and according to the reaction:



This is the reaction typically catalyzed by an anode, for example in a fuel cell, in the presence of hydrogen (in this case, the hydrogen is the fuel). The area under the peaks labeled HOR in FIG. 6 (active area  $A_{\text{active}}$ ) is directly proportional to the amount of charge (electrons) generated by the reaction and catalyzed by the Pt sites at the anode, and therefore is a good figure of merit for the anodic performance of the composition of the present invention. The larger the area under the peaks (e.g. HOR), the more charge (or electrons) generated, and the better the performance as an anode. More specifically, the active area represents charge resulting from the underpotential deposition of hydrogen ( $\text{H}_{\text{upd}}$ ) onto the composition acting as an anode.

Table 1 lists the active areas for various compositions of the present invention, as calculated by integrating the curves,

between 0 and 0.4 V vs. NHE, in FIG. 6, and dividing by the geometric surface area of the film being used as an electrode.

TABLE 1

Composition	Sample	Charge for $\text{H}_{\text{upd}}$ region	
		Active Area ( $A_{\text{active}}$ )	Film thickness (nm)
$\text{Pt}_{46}\text{Co}_{24}\text{Zr}_{20}$	E-11	397.66 $\mu\text{C}/\text{cm}^2$	88.3
$\text{Pt}_{68}\text{Co}_{23}\text{Zr}_9$	E-13	550.32 $\mu\text{C}/\text{cm}^2$	105.9
$\text{Pt}_{66}\text{Co}_{24}\text{Zr}_{10}$	E-23	741.64 $\mu\text{C}/\text{cm}^2$	137.8
$\text{Pt}_{69}\text{Co}_{20}\text{Zr}_{11}$	E-33	343.41 $\mu\text{C}/\text{cm}^2$	123.7
$\text{Pt}_{77}\text{Co}_{23}$	E-14	367.63 $\mu\text{C}/\text{cm}^2$	160.4

$A_{\text{active}}$  for a Pt film containing only Pt (pure Pt) is 210  $\mu\text{C}/\text{cm}^2$ . Thus, the data in Table 1 and FIG. 6 illustrate the films of the present invention have greater  $A_{\text{active}}$  as compared to a pure Pt film. Moreover, FIG. 6 shows the remarkable result that reducing the amount of Pt in the alloy electrode may increase the electrochemical performance of the Pt alloy electrode.

#### Durability (Non-Corrosive Properties)

FIG. 7 is a durability plot for a  $\text{Pt}_{67}\text{Zr}_{33}$  film, illustrating that compositions of the present invention do not corrode (to within the measurement accuracy) over a period of more than 1000 cycles. FIG. 7 compares the voltammogram taken after the film has been electrochemically prepared (by e.g., by cycling voltage over 100 cycles), with the voltammogram taken after 1000 cycles. The voltammograms are identical, to within the measurement accuracy of the experiment.

The durability of these compositions may exceed that of pure Pt, as the arrays have been extensively cycled ( $n > 10^3$  cycles) over the potential range 0.0 to 1.2 V, with no degradation of the electrode surface or decrease in electrochemical performance observed. Although not reported, representative electrodes of the Pt—Co—Zr thin films prepared have been cycled over the potential range 0.6-1.2 V, at a scan rate of 200 mV/sec, for at least 3000 cycles with no degradation in performance. These results suggest that alloys in Pt—Co—Zr composition manifold can be as active as Pt catalysts, although with much reduced Pt-loadings; i.e., with just ~50% of the amount of platinum metal.

#### Cathodic Behavior

The films of the present invention may be used as a cathode in the presence of, e.g., oxygen, to catalyze the reduction of the oxygen in an oxygen reduction reaction (ORR) (reacting the oxygen with, e.g., hydrogen ions produced at the anode).

For example, FIG. 8 illustrates the  $(\text{Pt}_3\text{Co})_{100-x}\text{Zr}_x$  thin film compositions synthesized are also electrochemically active for the ORR, a key point for application as a cathode material in an advanced fuel cell. FIG. 8 shows the results of potentiostatic polarization measurements conducted in a fully oxygenated 0.1 M  $\text{HClO}_4/\text{H}_2\text{O}$  electrolyte, using the set up of FIGS. 5a and 5b.

The ORR current is measured as a function of voltage applied to the film (voltage with respect to a normal hydrogen electrode defining 0 volts), in the presence of the oxygen in the electrolyte, wherein the voltage is swept at 1 mV/sec, from 1.05 V to 0.25 V (cathodic sweep). Large values for the ORR current density ( $\mu\text{A}/\text{cm}^2$ ) at 0.9 V (vs. NHE), is indicative of better cathodic performance. The cell current is divided by the geometric surface area of the film being used as an electrode to obtain current density. The data shown are for the ORR kinetic region of the polarization measurement and are taken after the surface of the films has become electrochemically active by performing 100 voltage cycles over 0 V to 1.2 V at a 200 mV/second scan rate. The ORR data were

quite reproducible for the alloys shown, with independent measurements providing nearly identical results.

FIG. 9 shows ORR current density for a voltage of 0.9 V vs NHE, for different alloy compositions of the present invention, wherein the ORR current density for a Pt thin film (Pt-TF) is also shown. The current density is normalized to the geometric area of the electrodes. FIG. 9 shows that the (111) oriented  $\text{Pt}_{66}\text{CO}_{24}\text{Zr}_{10}$  surface is ~30 times more active than the (111) Pt film measured in the same cell. The (111)  $\text{Pt}_{53}\text{CO}_{20}\text{Zr}_{27}$  surface is still ~17 times more active than Pt. The  $\text{Pt}_3\text{Co}$  current densities in this array, ~83  $\mu\text{Amps}/\text{cm}^2$ , are ~22 times greater than (111) Pt, in agreement with the literature. This behavior may be related to the intraalloy electron transfer between Pt/Co and Zr, and the d-band filling with x in these new ternary alloys.

While FIG. 9 shows that Pt—Co—Zr compositions with approximately 10% Zr are the best performers, given the high cost of Pt (~\$2000 per ounce currently), for some applications it may be advantageous to use less Pt, such as  $\text{Pt}_{53}\text{CO}_{20}\text{Zr}_{27}$  (sample E-31 with 53% Pt content), which still shows approximately 17 times enhancement over a pure Pt thin film cathode. The present invention illustrates that a cost vs. performance trade-off may be considered depending on the application.

#### Binary Compound

In a parallel study, the present invention has examined the properties of specific compositions in the binary alloy series,  $\text{Pt}_{100-x}\text{Zr}_x$ ,  $8 < x < 33$  (At. %). The present invention has synthesized two compositions, in thin film form, via co-sputtering. Both chemical compositions; e.g., 1)  $x=8$ ; and 2)  $x=33$ , illustrate the great potential of alloys in this binary series. As with the Pt—Co—Zr thin film compositions prepared, the  $\text{Pt}_{100-x}\text{Zr}_x$  alloys are stable in 0.1 M Perchloric acid and are electrochemically active for the ORR, with kinetic currents at 0.9 V that exceed elemental Pt. The 0.9 V (vs. NHE) enhancements of the ORR geometric current densities for  $\text{Pt}_{100-x}\text{Zr}_x$  are ~2.5 $\times$  greater for  $x=33$ , and 5.5 $\times$  greater for  $x=8$ , respectively.

FIG. 10 is a schematic illustrating the atomic positions of the Pt and Co in the  $\text{Pt}_3\text{Co}$  crystal. Also shown in FIG. 10 is the (111) plane.

The chemical stability for the binary  $\text{Pt}_{100-x}\text{Zr}_x$  alloys (with  $x=8$  and  $x=33$ ) is manifest in cyclic voltammograms much like those shown in FIG. 6, and for  $x=33$  in FIG. 7.

The present invention has also synthesized a  $\text{Pt}_3\text{Co}$  (sample E-13). The  $\text{Pt}_3\text{Co}$  sample also shows a good ORR kinetic current that is approximately 24 times greater than the ORR of a pure Pt film of comparable thickness, at 0.9 V vs. NHE.

FIG. 11a is a graph showing the binary phase diagram for the Pt—Zr alloy series, at which various  $\text{Pt}_{100-x}\text{Zr}_x$  compositions occur, wherein the ellipse shows compositions with high electrochemical performance, and FIG. 11b is a schematic showing the  $\text{D0}_{24}$  crystal structure for the  $\text{Pt}_3\text{Zr}$  composition.

#### Process Steps

FIG. 12 illustrates a method of fabricating a platinum (Pt) metal containing alloy composition useful as a catalyst electrode.

Block 1200 represents selecting a substrate. The step may comprise selecting a crystalline structure and grain size of the substrate. The substrate may be crystalline or nanocrystalline. The substrate may be a nanoparticle support. The substrate may be a hollow spherical nanoparticle (e.g., nanoshell) support, with typical diameter of 100 nm or larger. The substrate may be a micron-sized metallic or non-metallic support (e.g., ~325 mesh). The substrate may be a porous support having open areas. The substrate may be a wire or wire-gauze support. The substrate may include a metal (e.g., Ti, Au) and a

current collector structure and the metal's grain size is on a nanoscale (e.g., 100 nanometers or less).

Block 1202 represents selecting early transition metal A, A and B, or A, B and PGM. The amount of A or A and B and Pt may be such that the film has a microstructure ranging between an amorphous and a nanocrystalline microstructure. A may be at least one valve metal element of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb), for example. The step may comprise selecting the amounts of early transition metal, valve metal A, and Pt such that the film has a microstructure ranging between an amorphous or nanocrystalline microstructure, or with a mixture of both phases.

Block 1204 represents depositing Pt and the at least one early transition metal element A, or A and B onto the substrate.

A nanoparticle wash may be applied to the open areas of the porous support, wherein the nanoparticle wash includes a compound of Pt and at least one early transition metal element A, from group IVB or VB of the Periodic table. The nanoparticle wash may be heat-treated or fired to promote adherence to the surface of the porous support.

The alloy may be prepared by sputtering onto a support (or substrate). The sputtering may be from single alloyed target onto a support (or substrate). The sputtering may be co-sputtering from multiple targets onto a support (or substrate). The alloy may be deposited onto a support of any type, geometry, or size by hollow core magnetron sputtering.

The alloy may be prepared by electron beam evaporation from multiple targets onto a support (or substrate).

The alloy may be prepared by electron beam evaporation from a single alloyed target onto a support (or substrate).

Discrete nanoparticle forms of the alloy may be prepared by vapor-phase condensation in a high-vacuum chamber.

The alloy may be deposited onto a support by Atomic-Layer-Deposition (ALD) processing.

The alloy may be prepared by mechanical alloying (ball milling).

The alloy may be prepared by gas-atomization processing to yield an alloy in powder form.

The alloy may be prepared by centrifugal atomization processing to yield an alloy in powder form.

The alloy may be prepared by rapid solidification rate processing to yield an alloy with a nanostructured or amorphous atomic arrangement or microstructure. For example, the alloy may be prepared by the melt-spinning process to yield a thin-foil geometry product, wherein the alloy has a nanostructured or amorphous atomic arrangement or microstructure. The alloy may be prepared by the melt-extraction process to yield a wire-geometry product, wherein the alloy has a nanostructured or amorphous atomic arrangement or microstructure.

The alloy may be prepared by wet chemistry techniques resulting in the formation of small particles or nanoparticles. For example, the alloy may be prepared by the co-precipitation wet chemistry technique resulting in the formation of nanoparticles. The alloy may be prepared by the colloidal synthesis wet chemistry technique resulting in the formation of nanoparticles. The alloy may be prepared by the any wet chemistry technique involving PARR bomb processing at high temperatures resulting in the formation of nanoparticles.

The step may further comprise depositing the Pt, A and at least one late transition metal B on the substrate, wherein A includes at least one of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb), and B includes at least one of Cobalt (Co), Nickel (Ni), and Iron (Fe), and with at least one PGM element of Palladium (Pd), Ruthenium (Ru),

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Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir), for example. Alternatively, A may include at least one of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb), and B may include at least one of Cobalt (Co), Nickel (Ni), and Iron (Fe), for example.

The step may further comprise depositing the Pt, A, B, and at least one platinum group metal (PGM), wherein the PGM element includes at least one of Palladium (Pd), Ruthenium (Ru), Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir).

Block 1206 represents the composition fabricated using the method, a platinum (Pt) metal containing alloy composition for use as a catalyst (e.g., electrode). The Pt metal containing alloy composition may comprise a film (e.g. continuous) including a compound of Pt and at least one early transition metal element A from group IVB or VB of the Periodic table. The Pt metal containing alloy composition may comprise a nanoparticle; e.g., with nominal particle size 10 nm or less, including a compound of Pt and at least one early transition metal element A from group IVB or VB of the Periodic table. The Pt metal containing alloy composition may comprise a metallic nanocrystalline cluster or quantum dot on a nanoparticle support, wherein the metallic nanocrystalline cluster or quantum dot includes a compound of Pt and at least one early transition metal element A from group IVB or VB of the Periodic table. The composition may comprise less than 50% At. % Pt in the compound.

A may comprise at least one valve metal element of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb).

The Pt-containing compound further comprises at least one late transition metal B, thereby forming a Pt—B-A alloy.

A may include at least one of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb), and B includes at least one of Cobalt (Co), Nickel (Ni), and Iron (Fe).

The Pt-containing compound may further comprise at least one platinum group metal (PGM), thereby forming a Pt—PGM-B-A alloy, wherein the PGM element includes at least one of Palladium (Pd), Ruthenium (Ru), Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir).

A may include at least one of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb), and B includes at least one of Cobalt (Co), Nickel (Ni), and Iron (Fe), and with at least one PGM element of Palladium (Pd), Ruthenium (Ru), Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir).

A may include at least one of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb), and B includes at least one of Cobalt (Co), Nickel (Ni), and Iron (Fe), and with at least one PGM element of Palladium (Pd), Ruthenium (Ru), Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir).

Examples of embodiments of the alloy composition include  $\text{Pt}_{66}\text{Co}_{24}\text{Zr}_{10}$  (Atomic %),  $\text{Pt}_{68}\text{Co}_{23}\text{Zr}_9$  (Atomic %),  $\text{Pt}_{53}\text{Co}_{20}\text{Zr}_{27}$  (Atomic %),  $\text{Pt}_{92}\text{Zr}_8$  (Atomic %) and  $\text{Pt}_{67}\text{Zr}_{33}$  (Atomic %).

The film may be a continuous film on a nanoparticle support, on a hollow spherical nanoparticle support, or on a micron-sized metallic or non-metallic support, or on a wire or wire-gauze support, for example.

The Pt-containing compound may further comprise at least one late transition metal B, thereby forming a Pt—B-A alloy. For example, the Pt-containing alloy composition may be  $(\text{Pt}_3\text{Co})_{100-y}\text{Zr}_y$ , with  $0 \leq y \leq 30$  (At. %) or  $(\text{Pt}_{100-x}\text{Co}_x)_{100-y}\text{Zr}_y$ , with  $0 \leq x \leq 80$  and  $0.5 \leq y \leq 60$ .

The Pt-containing compound may further comprise at least one platinum group metal (PGM), thereby forming a Pt—PGM-B-A alloy, wherein the PGM element includes at

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least one of Palladium (Pd), Ruthenium (Ru), Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir).

The film may be crystalline with a grain size such that, when the film is used as the cathode catalyst, an active area of the cathode is above  $218 \mu\text{C}/\text{cm}^2$  (see also FIG. 4a, FIG. 6, and Table 1, for example). The film may be sputtered on the substrate.

The alloy composition may be crystalline (e.g., with a (111) crystallographic orientation), nanocrystalline with a grain size of no more than 100 nm, nanocrystalline with a preferred grain size of less than 10 nm.

Block 1208 represents using the composition formed in Block 1206 as a catalyst.

The alloy may be used as a catalyst in a catalytic convertor for an internal combustion engine burning gasoline or diesel fuel.

The alloy may be used as a replacement catalyst for Pt catalysts used in the Platforming process, which enables the synthesis of gasoline without the addition of lead to the gasoline.

The alloy may be used as the anode in a hydrogen-air fuel cell or hydrogen-oxygen fuel cell.

The alloy may be used as the cathode in a hydrogen-air fuel cell, hydrogen-oxygen fuel cell, or direct methanol fuel cell (DMFC). In this application, the amount of A and an amount of Pt in the alloy may be such that the alloy may be at least 2 times more electrochemically active, than Pt, as the cathode material for the oxygen reduction reaction (ORR).

The alloy may be used as the anode in a direct methanol fuel cell (DMFC).

In one embodiment, the alloy composition is  $\text{Pt}_{66}\text{Co}_{24}\text{Zr}_{10}$  (Atomic %), and may be used as the cathode in a hydrogen-air fuel cell. In this embodiment, the alloy as a cathode may be at least 31 times more electrochemically active at 0.9 V (vs. NHE) than Pt, when used as the cathode material for the oxygen reduction reaction (ORR).

In one embodiment, the alloy composition is  $\text{Pt}_{68}\text{Co}_{23}\text{Zr}_9$  (Atomic %), and may be used as the cathode in a hydrogen-air fuel cell. In this embodiment, the alloy as a cathode may be at least 30 times more electrochemically active at 0.9 V (vs. NHE) than Pt, when used as the cathode material for the oxygen reduction reaction (ORR).

In another embodiment, the alloy composition is  $\text{Pt}_{53}\text{Co}_{20}\text{Zr}_{27}$  (Atomic %), and may be used as the cathode in a hydrogen-air fuel cell. In this example, the alloy cathode may be at least 16 times more electrochemically active at 0.9 V (vs. NHE) than Pt, when used as the cathode material for the oxygen reduction reaction (ORR).

In another embodiment, the alloy composition is  $\text{Pt}_{92}\text{Zr}_8$  (Atomic %), and may be used as the cathode in a hydrogen-air fuel cell. In this example, the alloy as a cathode is at least 5 times more electrochemically active at 0.9 V (vs. NHE) than Pt, when used as the cathode material for the oxygen reduction reaction (ORR).

In another embodiment, the alloy composition is  $\text{Pt}_{67}\text{Zr}_{33}$  (Atomic %), and may be used as the cathode in a hydrogen-air fuel cell. In this example, the alloy cathode may be at least 2 times more electrochemically active at 0.9 V (vs. NHE) than Pt, when used as the cathode material for the oxygen reduction reaction (ORR).

An amount of A and the amount of Pt in the film may be such that the film is electrochemically stable, with no decrease in electrochemical activity of the cathode, when the electrochemical activity is characterized by conducting multiple cycles of a cyclic voltammogram (CV) over the potential range 0.0 to at least 1.2 Volts (vs. NHE) to determine the durability of said alloy: measuring a current flowing between

the film and an anode, in a electrochemical  $\frac{1}{2}$ -cell, in response to a voltage applied to the film that is ramped from 0.0 to at least 1.2 Volts at a scan rate of at least 10 mV/second, wherein the voltage is ramped from the 0.0 Volts to the at least 1.2 Volts at least 1000 times (over 1000 cycles); and the current not varying by more than the measurement accuracy (+/-10 microamps).

When the alloy is used as the cathode or anode in a hydrogen-air fuel cell, hydrogen-oxygen fuel cell, or direct methanol fuel cell (DMFC), the electrolyte may comprise an acid electrolyte, the acid electrolyte may comprise a mixture perchloric acid in water ( $\text{HClO}_4/\text{H}_2\text{O}$ ) (e.g., a 1 molar perchloric acid concentration in water ( $\text{HClO}_4/\text{H}_2\text{O}$ )), the acid electrolyte may comprise a mixture of sulfuric acid in water ( $\text{H}_2\text{SO}_4/\text{H}_2\text{O}$ ) (e.g., a 1 molar sulfuric acid concentration in water ( $\text{H}_2\text{SO}_4/\text{H}_2\text{O}$ )). The acid electrolyte may saturate a polymer exchange membrane electrolyte (e.g. Nafion).

The film may be crystalline with a grain size such that, when the film is used as the cathode catalyst in a hydrogen-air or hydrogen-oxygen fuel cell, and the measured Pt-active area from the hydrogen-oxidation-reaction (HOR) of the cathode is greater than that of polycrystalline Pt,  $A_{\text{Pt}}=210 \mu\text{C}/\text{cm}^2$ .

#### Possible Modifications

The compositions of the present invention may be fabricated by methods other than co-sputtering, including for example, e-beam deposition.

The compositions of the present invention may be used for various applications where catalytic properties are useful. For example:

1. The present invention may be used to replace Pt gauze in a nitrogen fertilizer application.

2. The present invention may be used in a catalytic converter.

3. The present invention may be used in a hydrogen fuel cell or an hydrogen and air fuel cell, or in a fuel cell that uses a PEM membrane electrode. The present invention may be used as an anode and/or a cathode.

4. The present invention could be used in a nanofuel cell, wherein the nanofuel cell is surrounded by microcatalytic compositions of the present invention.

5. Various crystal forms of the present invention, as well as various compositions, may be tailored for particular applications.

#### Advantages and Improvements

Technical issues in the current state of the art, coupled with the high-cost and limited availability of Pt metal, have motivated the present invention to search for new Pt-based, transition metal alloy catalysts that are stable in acid and electrochemically active for the oxygen reduction reaction (ORR).

The scientific methodology of the present invention has employed the following key concepts in materials design, chemical physics, and electrochemistry, to achieve improvements associated with the following metrics for fuel cell cathode materials:

1) controlled alloy design methods used to obtain improved corrosion resistance and increased stability in acid solutions;

2) thin film synthesis (co-sputtering unique) used to obtain a wide range of multi-component Pt-based, binary- and ternary-alloys, thin film form;

3) thin film synthesis, coupled with modulations of chemical composition, enables control of the microstructural length scale of the materials prepared (amorphous to nanocrystalline microstructures);

4) d-band engineering (i.e., filling of the transition metal d-band) used to control the relative position of the Fermi energy,  $\epsilon_f$ , and the density of states at the Fermi energy,

$\mathcal{D}(\epsilon_f)$ , thereby enabling controlled reductions in the ORR activation potential (i.e., reduced over-potential  $V_{\text{over}}$  results in increased ORR current density);

5) control of the microstructural length scales can be used to exploit the occurrence of quantum size effects in small- or nanocrystalline-particles; which can result in controllable shifts in the relative position of the Fermi energy,  $\epsilon_f$ , thereby enabling another means to obtain reductions in the ORR activation potential.

The results of the present invention suggest that in order to endow an enhanced corrosion resistance to Pt-based alloys, while maintaining a high catalytic activity, addition of group IVB, VB valve metal elements appears to be of value. The present invention is also able to achieve at least 2-10 times electrochemical (ORR and Hydrogen oxidation) performance as compared to a pure Pt electrode. The present invention also shows that reducing the amount of Pt in the alloy electrode may increase the electrochemical performance of the Pt alloy electrode, as shown in FIGS. 6-9.

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#### CONCLUSION

This concludes the description of the preferred embodiment of the present invention. The foregoing description of one or more embodiments of the invention has been presented for the purposes of illustration and description. It is not intended to be exhaustive or to limit the invention to the

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precise form disclosed. Many modifications and variations are possible in light of the above teaching. It is intended that the scope of the invention be limited not by this detailed description, but rather by the claims appended hereto.

What is claimed is:

1. A platinum (Pt) metal containing alloy composition, comprising:

a compound including an alloy of Pt and at least one valve metal element A from group IVB or VB of the Periodic table having

an amount of the valve metal; and

a structure of the alloy that is single phase, wherein the alloy's surface exhibits a catalytic reaction and the alloy's durability is at least as high as:

(i) pure Pt exhibiting the catalytic reaction, or

(ii) a catalyst consisting essentially of Pt exhibiting the catalytic reaction.

2. The composition of claim 1, wherein the alloy is a film and the film is a continuous film on a nanoparticle support.

3. The composition of claim 1, wherein the alloy is a film and the film is a continuous film on a hollow spherical nanoparticle or nanoshell support.

4. The composition of claim 1, wherein the alloy is a film and the film is a continuous film on a micron-sized metallic or non-metallic support.

5. The composition of claim 1, wherein the alloy is a film and the film is a continuous film on a wire or wire-gauze support.

6. The composition of claim 1, having less than 50% Atomic % of Pt in the alloy.

7. The composition of claim 1, wherein valve metal element A is at least one of Zirconium (Zr), Titanium (Ti), Hafnium (Hf), and Niobium (Nb).

8. The composition of claim 7, wherein the Pt-containing compound further comprises at least one transition metal B, thereby forming a Pt—B-A alloy, and wherein B includes at least one of Cobalt (Co), Nickel (Ni), and Iron (Fe).

9. The composition of claim 8, wherein the Pt-containing compound further comprises at least one platinum group metal (PGM), thereby forming a Pt—PGM-B-A alloy, wherein the PGM element includes at least one of Palladium (Pd), Ruthenium (Ru), Rhodium (Rh), Rhenium (Re), Osmium (Os), and Iridium (Ir).

10. The composition of claim 8, wherein the alloy is  $(Pt_3Co)_{100-y}Zr_y$ , with  $0 \leq y \leq 30$  and y is Atomic %.

11. The composition of claim 8, wherein the alloy is  $(Pt_{100-x}Co_x)_{100-y}Zr_y$ , with  $0 \leq x \leq 80$  and  $0.5 \leq y \leq 60$ .

12. The composition of claim 1, wherein the alloy is crystalline.

13. The composition of claim 12, wherein the alloy is nanocrystalline with a grain size of no more than 100 nm.

14. The composition of claim 13, wherein the alloy is nanocrystalline with a grain size of less than 10 nm.

15. The composition of claim 12, wherein the alloy has a (111) crystallographic orientation.

16. The composition of claim 12, wherein the alloy is a film and an amount of A and the amount of Pt in the film is such that the film is electrochemically stable, with no decrease in electrochemical activity of the film, when the electrochemical activity is characterized by conducting multiple cycles of a cyclic voltammogram (CV) over a potential range of 0.0 to at least 1.2 Volts (vs. NHE), to determine the durability of said alloy:

measuring a current flowing between the film and an anode, in an electrochemical  $\frac{1}{2}$  cell, in response to a voltage applied to the film that is ramped from 0.0 to at least 1.2 Volts at a scan rate of at least 10 mV/second,

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wherein the voltage is ramped from the 0.0 Volts to the at least 1.2 Volts at least 1000 times (over 1000 cycles); and the current not varying by more than the measuring's accuracy ( $\pm 10$  microamps).

17. The composition of claim 1, wherein the alloy is a cathode in a hydrogen-air fuel cell (H2-Air FC), hydrogen-oxygen fuel cell (H2-O2 FC), or an anode in a direct methanol fuel cell (DMFC), and

an amount of A and an amount of Pt in the alloy is such that that the H2-Air FC or H2-O2 FC cathode is at least 2 times more electrochemically active for an oxygen reduction reaction (ORR), than a cathode consisting essentially of Pt,

an amount of A and an amount of Pt in the alloy is such that that the H2-Air FC or H2-O2 FC anode is more electrochemically active for the hydrogen oxidation reaction (HOR), than an anode consisting essentially of Pt, or

an amount of A and an amount of Pt in the alloy is such that that the DMFC anode is more electrochemically active for the methanol oxidation reaction (MOR), than an anode consisting essentially of Pt.

18. The composition of claim 1, wherein the alloy is a film and the film is crystalline with a grain size such that, when the film is a catalyst for an anode in a hydrogen-air or hydrogen-oxygen fuel cell, a measured Pt-active area ( $A_{Pt}$ ) from a hydrogen-oxidation-reaction (HOR) of the anode is greater than that of polycrystalline Pt, or greater than  $A_{Pt} = 210 \mu C/cm^2$ .

19. The composition of claim 1, wherein the alloy is a film and amounts of 1 A and Pt are such that the film has a microstructure ranging between a amorphous or nanocrystalline microstructure, or with mixture of both phases.

20. The composition of claim 1, further comprising a nanoparticle including the alloy.

21. The composition of claim 1, further comprising a metallic nanocrystalline cluster or quantum dot on a nanoparticle support, wherein the metallic nanocrystalline cluster or quantum dot includes the alloy.

22. A method of fabricating a platinum (Pt) metal containing alloy composition, comprising:

fabricating a compound including an alloy of Pt and at least one valve metal element A, to obtain an amount of the valve metal and a single phase structure of the alloy, wherein the alloy's surface exhibits a catalytic reaction and the alloy's durability is at least as high as:

(i) pure Pt, or

(ii) a catalyst consisting essentially of Pt.

23. The method of claim 22, wherein the alloy is fabricated by evaporation from multiple targets or from a single alloyed target onto a support (or substrate).

24. The composition of claim 1, wherein:

the alloy is disposed within a catalytic converter system for vehicle emissions,

the alloy exhibits catalytic properties,

the catalytic properties break down harmful byproducts present in hot exhaust streams of the vehicle emissions, including performing one or more of the following reactions:

1) reduction of nitrous oxides to nitrogen and oxygen;

2) oxidation of carbon monoxide to carbon dioxide; and

3) oxidation of unburned hydrocarbon (HCs).

25. A catalytic converter or fuel cell apparatus comprising: an alloy of Pt and at least one valve metal element A from group IVB or VB of the Periodic table having

an amount of the valve metal and

a structure of the alloy that is single phase, wherein the alloy's surface exhibits a catalytic reaction in the cata-

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lytic converter or fuel cell apparatus and the alloy's durability and catalytic activity is at least as high as:

- (i) pure Pt exhibiting the catalytic reaction, or
- (ii) a catalyst consisting essentially of Pt exhibiting the catalytic reaction.

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**26.** The apparatus of claim **25**, wherein the apparatus is a catalytic converter and the catalytic reaction breaks down byproducts present in emissions received in the catalytic converter.

**27.** The apparatus of claim **25**, wherein the apparatus is a hydrogen-air fuel cell (H<sub>2</sub>-Air FC), hydrogen-oxygen fuel cell (H<sub>2</sub>-O<sub>2</sub> FC), or a direct methanol fuel cell (DMFC).

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